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Perspective on the hyperpolarisation technique Signal Amplification By Reversible Exchange (SABRE) in NMR spectroscopy and MR imaging

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Abstract

Signal Amplification By Reversible Exchange (SABRE) is a *parahydrogen* based technique that utilises a metal complex, normally centred on iridium, to propagate polarisation from *parahydrogen* derived hydride ligands to spin-½ nuclei located in a bound substrate. To date, substrates possessing ¹H, ¹³C, ¹⁵N, ¹⁹F, ³¹P, ²⁹Si and ¹¹⁹Sn nuclei have been polarised by this technique. The exact positioning of these nuclei has a direct bearing on the enhancement observed and so substrates must be chosen or synthesised with care in order to maximise polarisation transfer, and hence the resulting enhancement. The chemical composition of the metal complex must be similarly appraised, as the exchange rate of substrates and *parahydrogen* are implicated heavily in efficient polarisation transfer. The nature of the polarisation transfer, whether homogenous or heterogeneous, is another important facet to consider here, as is conducting SABRE in water based systems. This review discusses the physical and theoretical aspects of the SABRE experiment, as well as the applications of the SABRE technique, namely the detection of analytes at concentrations far below what would be possible with conventional NMR techniques and the collection of hyperpolarised magnetic resonance images. Advances relating to utilising singlet states for SABRE, pulse sequence design and the nature of the polarisation transfer mechanism are also discussed and the implications for future SABRE-based discoveries highlighted.

Keywords : Hyperpolarisation, SABRE, *parahydrogen*, NMR, MRI

Introduction

Nuclear Magnetic Resonance (NMR) is an invaluable tool for ascertaining the chemical structures of materials and for gaining insight into rates of reaction, for example. Similarly, Magnetic Resonance Imaging (MRI) is used routinely for the collection of anatomical images of the human body, which can be used to improve diagnosis and treatment strategies. However, both techniques suffer from an inherent sensitivity problem. Thermal energy effects, which are governed by $k_B T$ (where k_B is the Boltzmann constant), are sufficiently large compared to the energy difference that exists between the two possible energy states that arise from the Zeeman splitting. Boltzmann statistics directly affect the population distribution of the states, and as such, broadly equally-populated states result. This population distribution is termed as polarisation (P) and can be calculated according to Equation 1. For example, at 298 K, ¹H nuclei exposed to a 9.4 T field are polarised to 3.2×10^{-5} (0.0032%) or a population distribution that differs by roughly one in every 31,000 nuclei.

$$P = \frac{\gamma h B_0}{2k_B T}$$

Equation 1. Expression for calculating thermal polarisation. γ is the gyromagnetic ratio, h is Planck's constant, B_0 is the magnetic field to which nuclei are exposed to, k_B is the Boltzmann constant and T is the temperature.

Due to the Boltzmann-like distribution of nuclear spins across energy levels, the signal intensity observed in the ^1H NMR spectrum and the ^1H MR image, for example, is only a fraction of its true potential. Furthermore, for nuclei such as ^{13}C that possesses a gyromagnetic ratio of one quarter of that of ^1H , the polarisation is significantly lower (Table 1). The inherent insensitive nature of NMR and MRI led to the development, and employment, of hyperpolarisation methods.

Table 1. Polarisation values for ^1H , ^{13}C , ^{19}F and ^{31}P nuclei at different magnetic fields and temperatures.

Nucleus	Polarisation at $B_0 = 1.5$ T; T = 310 K	Polarisation at $B_0 = 1.4$ T; T = 298 K	Polarisation at $B_0 = 9.4$ T; T = 298 K
^1H	4.9×10^{-6}	4.8×10^{-6}	3.2×10^{-5}
^{13}C	1.2×10^{-6}	1.2×10^{-6}	8.1×10^{-6}
^{15}N	5.0×10^{-7}	4.9×10^{-7}	3.3×10^{-6}
^{19}F	4.7×10^{-6}	4.5×10^{-6}	3.0×10^{-5}
^{31}P	2.0×10^{-6}	1.9×10^{-6}	1.3×10^{-5}

The simplest form of hyperpolarisation is brute force [1]. Brute force techniques centre on manipulating Equation 1, in particular the temperature and magnetic field strength, to improve polarisation. Increasing the magnetic field in which the measurement takes place and / or decreasing the temperature of the sample being analysed results in better polarisation. However, these are only small gains. For example, for ^1H nuclei at 4 K and 9.4 T the polarisation is 2.4×10^{-3} , which is two orders of magnitude larger than at 298 K. Increasing the magnetic field strength to 20 T only increases polarisation by a further *ca.* 50% ($P = 5.11 \times 10^{-3}$). Achieving both a stronger magnetic field and / or low temperature is costly and in some cases, very impractical (e.g. MR images of humans cannot be acquired at 4 K). Despite this, successful studies have been conducted, such as the hyperpolarisation of ^{13}C -labelled pyruvic acid, which was enhanced by 1600-fold [2]. However, brute force techniques have been largely superseded by other hyperpolarisation techniques, namely DNP (Dynamic Nuclear Polarisation) [3-7], SEOP (Spin Exchange Optical Pumping) [8] and PHIP (Parahydrogen Induced Polarisation) [9-14]. SEOP and DNP fall out of the scope of this review and hence readers are directed to the references indicated for information relating to these techniques.

PHIP utilises *parahydrogen*, a nuclear singlet, as the source of polarisation. At room temperature, H_2 gas consists of 25% *parahydrogen* and 75% *ortho*hydrogen. The interconversion between the two forms is spin forbidden and as such, a catalyst must be used to relax the spin selection rules. Fe_2O_3 [15], charcoal [10] and silica/ FeCl_3 [11] fulfil this role very well and have been utilised extensively in the literature. In the absence of catalyst the conversion process can take years to occur. The H_2 gas also needs to be cooled so that *parahydrogen* is produced in larger amounts relative to *ortho*hydrogen. At 77 K, the ratio is 50 : 50 *ortho*hydrogen to *parahydrogen* whereas at 20.3 K (boiling point of H_2), 99.859% *parahydrogen* is produced [12]. The requirement to cool the H_2 gas becomes apparent when the wavefunction of H_2 is considered. For the inversion of the two protons (fermions), only the rotational and nuclear spin components of the wavefunction change. The wavefunction has to be antisymmetric with respect to inversion. Thus, *parahydrogen* is associated with symmetric rotational values (even values of J) whereas *ortho*hydrogen has asymmetric rotational values (odd values of J). For the former $S(S-1)/2$ nuclear spin states exist whereas the latter has $S(S+1)/2$ nuclear spin states (where $S=2s+1$ ($s = \frac{1}{2}$)). This equates to one and three states respectively, and therefore, *parahydrogen* exists in a nuclear singlet state whereas *ortho*hydrogen exists in a nuclear triplet state. In the presence of a catalyst that relaxes the spin selection rules, improved conversion of *ortho*hydrogen to *parahydrogen* therefore results at lower temperatures as the $J = 0$ becomes more populated and is 1.4 kJ mol^{-1} lower in energy than the $J = 1$ state. The

production of *parahydrogen* has been automated in order to meet the specific requirements of some groups [15-17].

Bowers and Weitekamp first postulated the use of *parahydrogen* to obtain very large nuclear-spin polarisations in 1986 [18]. They predicted that following chemical reaction, and subsequent radio-frequency (RF) irradiation, the scalar *parahydrogen* state would be converted into polarisation of order unity on the nuclear spins of the hydrogenation products formed. It is worth noting that *parahydrogen* has a total spin of zero and is not visible by NMR and thus its symmetry must be broken in order to observe the polarised state. Furthermore, they stated that the phenomena should have “wide utility in the study of chemical reactions of molecules by hydrogen addition or subsequent reactions”. The following year they experimentally proved it by hydrogenating acrylonitrile to propionitrile using Wilkinson’s catalyst ($[\text{RhCl}(\text{PPh}_3)_3]$) using *parahydrogen* within the magnetic field used for measurement [19]. The ^1H NMR spectrum showed large anti-phase multiplets for the *parahydrogenated* product following the application of a $\pi/4$ pulse. The title of the paper “*Parahydrogen and Synthesis Allow Dramatically Enhanced Nuclear Alignment*”, which was abbreviated to PASADENA, became synonymous with PHIP conducted within the magnetic field. Around the time of this publication, a similar observation had been made by the group of Eisenberg but had been mistakenly interpreted as spin sorting of a radical pair intermediate (otherwise known as chemically induced dynamic nuclear polarisation (CIDNP)) [20]. The storage of the sample in liquid nitrogen prior to reaction at room temperature had most likely created a sufficient concentration of *parahydrogen* to detect polarised signals [9]. In a further extension, the *parahydrogenation* was conducted at low-field and then the sample rapidly transported (adiabatically) to the measurement field for detection of the hyperpolarised signal. This approach is, therefore, different to PASADENA and thus the acronym ALTADENA (Adiabatic Longitudinal Transport After Dissociation Engenders Nuclear Alignment) was coined to differentiate that the polarised material was produced at low-field rather than at high-field [21]. The *parahydrogenation* of styrene using Wilkinson’s Catalyst was the first substrate to be polarised by ALTADENA. A $\pi/2$ pulse is required to optimally read-out the magnetisation created, and the signals present as in-phase signals, although they can be emissive or absorptive. The acronym PHIP [9] is used to broadly describe ALTADENA and PASADENA type reactions.

For the purposes of this work, SABRE (Signal Amplification By Reversible Exchange), will be the focus. Developments and applications of the technique have been previously reviewed [22]. Unlike other PHIP techniques, polarisation is transferred without chemical modification of the analyte. First reported in 2009, SABRE relies upon transferring polarisation from *parahydrogen* derived hydrides to the spin- $\frac{1}{2}$ nuclei of ligated ligands through the *J*-coupling network at low magnetic field [23]. A metal catalyst is used to bring the two entities in to contact with one another. Both the ligand and hydrides can dissociate from the catalyst allowing other molecules to bind, and as such, the process can be repeated many times. Figure 1 exemplifies the process. Practically, this involves dissolving the catalyst and substrate in solvent and then either bubbling *parahydrogen* through the solution, or using a Young’s capped NMR tube with an atmosphere of *parahydrogen* and physically shaking it. This leads to a build-up of polarisation within analyte molecules not ligated to the metal complex, which can be read-out by the application of RF pulses within the measurement field, following rapid transfer from low- to high-magnetic field. The precise nature of the polarisation transfer process is dependent on both scalar coupling and chemical shift effects [24]. Thus, polarisation is both magnetic field- and time-dependent, the latter reflecting the transient nature of the complex to which *parahydrogen* derived hydrides and substrate ligands are bound. Theoretical modelling has shown that the substrate ^1H nuclear spins, following polarisation transfer, evolve into states, that when interrogated by RF pulses, produce in-phase signals originating from longitudinal

magnetisation and antiphase signals originating from either longitudinal two-spin order or residual zero quantum terms. Under conditions where the chemical shift evolution is suppressed by locating the sample in a zero-field or through the substrate possessing isochronous spins, polarisation is predominantly transferred through spins and thus leads to the generation of a singlet state. This has been exemplified for a number of systems and is discussed further in the section entitled “Generation and utilisation of long-lived states by SABRE”.

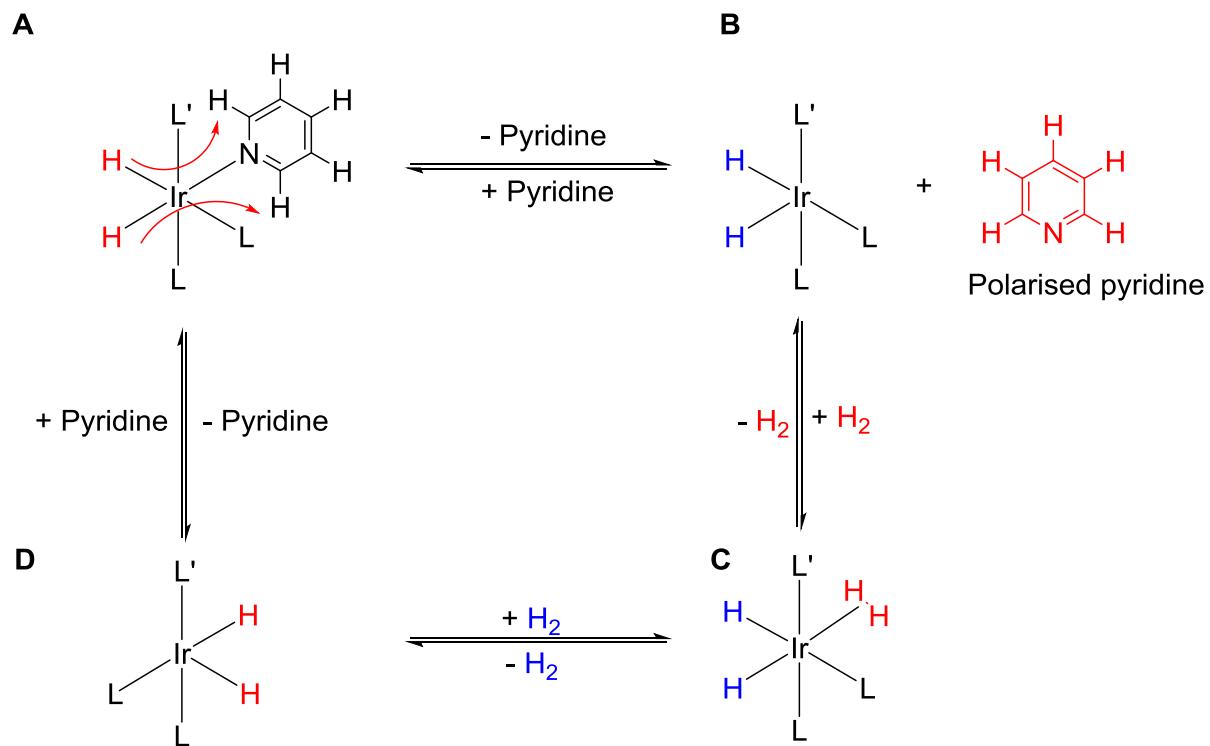


Figure 1. Mechanism for the transfer of polarisation to pyridine (L). L' represents a ligand not involved in the transfer mechanism e.g. a carbene or phosphine ligand. A) Transfer of polarisation through the *J*-coupling from *parahydrogen* derived hydride ligands to spin-½ nuclei of pyridine. B) Hyperpolarised pyridine dissociates leaving a vacant site. C). *Parahydrogen* occupies the vacant site. D) Hydrogen exchange resulting in *parahydrogen* derived hydrides reforming and loss of *ortho*hydrogen.

Early observations of the SABRE effect

SABRE was first exemplified using cationic iridium complexes of the type $[\text{Ir}(\text{COD})(\text{PR}_3)_2]\text{BF}_4$, where R = Ph, *p*-tolyl or *p*-C₆H₄-OMe) [25]. These complexes are essentially pre-catalysts as they not capable of polarisation transfer. In the presence of pyridine and *parahydrogen* these complexes were shown to form *trans,cis,cis*-[Ir(PR₃)₂(py)₂(H)₂]⁺ (py = pyridine) and small amounts of *fac,cis*-[Ir(PR₃)(py)₃(H)₂]⁺; these are the active polarisation transfer catalysts. These complexes are readily observable in the ¹H NMR due to the differences in the chemical shift of their hydride ligands. For example, when R = Ph, the hydride of *trans,cis,cis*-[Ir(PPh₃)₂(py)₂(H)₂]⁺ occurs at δ -21.64 whereas the hydride of *fac,cis*-[Ir(PPh₃)(py)₃(H)₂]⁺ is observed at δ -21.8. When ¹⁵N-labelled pyridine and a *parahydrogen* atmosphere were employed, both hydride signals are polarised in the ¹H NMR spectrum.

Interestingly, the hydride signal of *fac,cis*-[Ir(PPh₃)(py)₃(H)₂]⁺ is much larger than the corresponding signal of *trans,cis,cis*-[Ir(PPh₃)₂(py)₂(H)₂]⁺ despite the latter being in substantially larger concentration (in the presence of a 9-fold excess of pyridine, relative hydride resonance intensities are 0.04 : 1 respectively). The propensity of these complexes to exchange H₂ in *trans,cis,cis*-[Ir(PPh₃)₂(py)₂(H)₂]⁺ was probed using normalized hydride signal intensity ratios as a guide. These data revealed that the

exchange process increases in the order $\text{P}(p\text{-C}_6\text{H}_4\text{-OMe})_3 > \text{P}(p\text{-tolyl})_3 > \text{PPh}_3$. This trend mirrors the increase in the electron donating ability of the phosphine. Conversely, the trend observed for the H_2 exchange in $\text{fac},\text{cis}-[\text{Ir}(\text{PPh}_3)(\text{py})_3(\text{H})_2]^+$ is the opposite and thus the extent of enhancement follows the reduction of electron donating power of the phosphine. It was noted, however, that EXSY (Exchange Spectroscopy) measurements revealed little or no exchange of hydride into free hydrogen at 335 K in these complexes, which is suggestive that reductive elimination of H_2 under these conditions is not facile. Pyridine exchange rates for $\text{trans},\text{cis},\text{cis}-[\text{Ir}(\text{PR}_3)_2(\text{py})_2(\text{H})_2]^+$ were determined and values of 0.4 s^{-1} , 0.82 s^{-1} and 0.90 s^{-1} were obtained at 335 K for when $\text{R} = \text{Ph}$, $p\text{-tolyl}$ or $p\text{-C}_6\text{H}_4\text{-OMe}$ respectively. Obtaining pyridine exchange rates at 335 K for $\text{fac},\text{cis}-[\text{Ir}(\text{PR}_3)(\text{py})_3(\text{H})_2]^+$ proved more difficult due to decomposition ($\text{R} = p\text{-C}_6\text{H}_4\text{-OMe}$), or interconversion of the axial and equatorial pyridine ligands ($\text{R} = \text{Ph}$). However, a rate of 2.0 s^{-1} was obtained for when $\text{R} = p\text{-tolyl}$; this is *ca.* 2.4 times quicker than the analogous *bis*-phosphine complex. Pyridine exchange rates for all three complexes were obtained at 308 K and were similar (*ca.* 0.07 s^{-1}) and larger than those of $\text{trans},\text{cis},\text{cis}-[\text{Ir}(\text{PR}_3)_2(\text{py})_2(\text{H})_2]^+$ complexes. Enthalpy data also suggested that Ir-N bond is strengthened in $\text{trans},\text{cis},\text{cis}-[\text{Ir}(\text{PR}_3)_2(\text{py})_2(\text{H})_2]^+$ complexes relative to $[\text{Ir}(\text{PR}_3)(\text{py})_3(\text{H})_2]^+$ complexes. Resultantly, the $[\text{Ir}(\text{PR}_3)(\text{py})_3(\text{H})_2]^+$ complexes represent the better class of complex to use for polarisation transfer in SABRE.

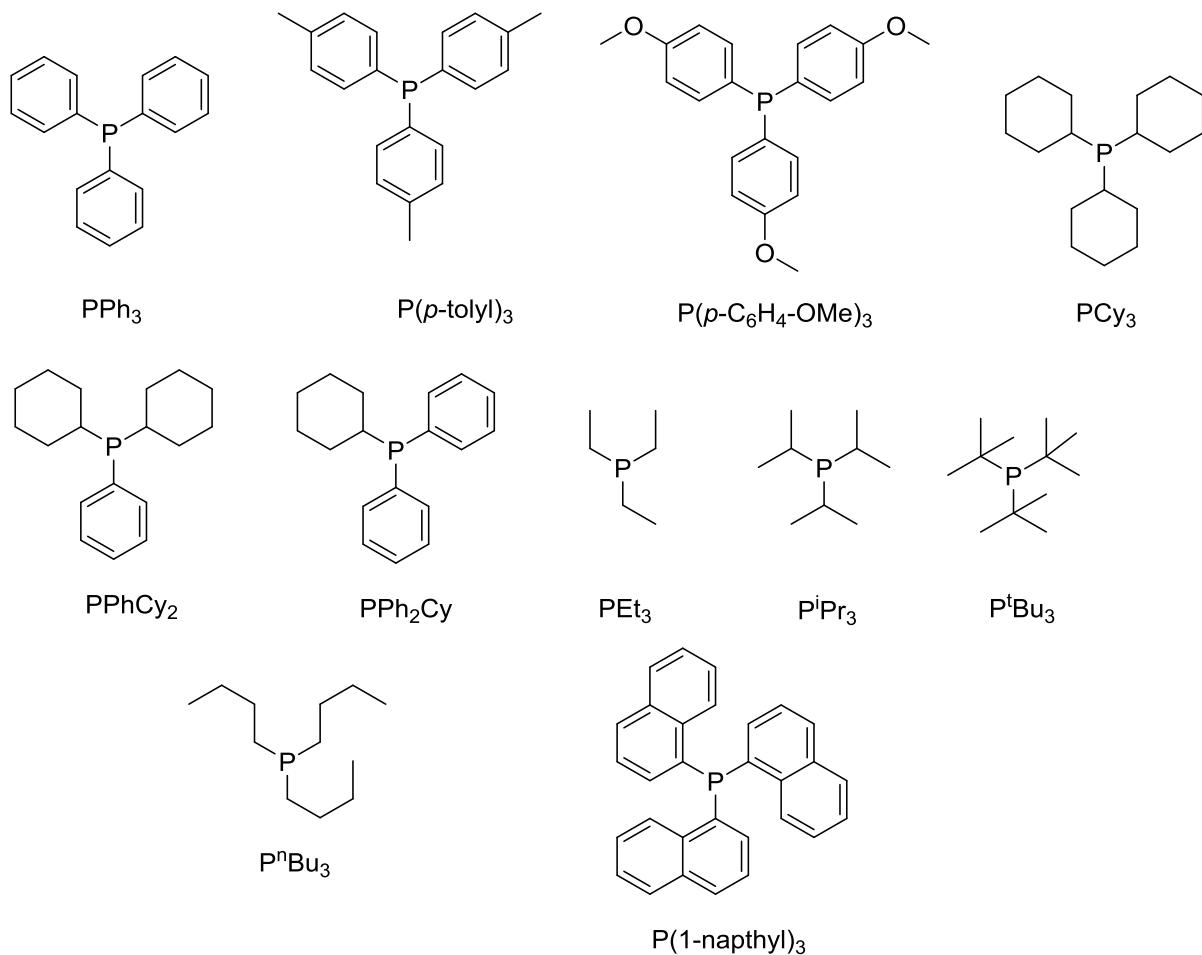


Figure 2. Phosphine ligands used in the construction of SABRE pre-catalysts.

The knowledge gleaned on the identity of the phosphine ligand led to a further study to investigate the use of $[\text{Ir}(\text{COD})(\text{PCy}_3)(\text{py})]\text{BF}_4$, Crabtree's catalyst, to polarise pyridine in the presence of *parahydrogen* [26]. The active form of the catalyst, based on NMR studies and X-ray crystallographic data, proved to be $\text{fac},\text{cis}-[\text{Ir}(\text{PCy}_3)(\text{py})_3(\text{H})_2]\text{BF}_4$. The ^1H NMR signals of free pyridine were shown to

be enhanced when a sample containing an excess of pyridine relative to the form of the catalyst was polarised in earth's magnetic field. The *ortho*-, *para*- and *meta*-¹H sites of pyridine were enhanced ~144-, ~47- and 220-fold respectively at 298 K. The enhancement is a measure of the increase in integral size in the hyperpolarised spectrum relative to the same integral in the non-hyperpolarised spectrum. Subsequent investigations revealed that polarisation transfer did not occur in the magnetic field used for measurement (9.4 T) or that the signals of pyridine were being enhanced through chemical incorporation of *parahydrogen*, perhaps through a C-H bond activation route. Polarisation transfer to ¹³C and ¹⁵N was also demonstrated. For example, the polarised ¹⁵N NMR spectrum displays a 128-fold increase in signal strength for pyridine relative to the thermal spectrum. Furthermore both free and bound (*trans* to hydride) pyridine ¹⁵N signals are evident (¹⁵N labelled pyridine was used in this study). Probing the identity of the phosphine yielded more results relating to the polarisation transfer process. It was observed that the more electron rich and most sterically demanding phosphines yielded larger relative enhancements. Thus, of the phosphines investigated (PCy₃, PPhCy₂, PPh₂Cy, PEt₃, P^tPr₃, PⁿBu₃, P^tBu₃ and P(1-naphthyl)₃), it was PPhCy₂, the most electron rich and sterically demanding phosphine, that yielded the most successful transfer of polarisation to pyridine.

These early studies demonstrated that sterically bulky electron-donating phosphines delivered the highest levels of transferred hyperpolarisation. Thus, exchanging phosphines for *N*-heterocyclic carbenes (NHCs) might be expected to deliver even better results. Thus, the complex [Ir(IMes)(COD)Cl] (IMes = 1,3-bis(2,4,6-trimethylphenyl)imidazole-2-ylidene, COD = cyclooctadiene) was synthesised and tested for SABRE activity [27]. In the presence of *parahydrogen* and pyridine, the complex [Ir(IMes)(py)₃(H)₂]⁺ is formed (Figure 3). The crystal structure of [Ir(IMes)(py)₃(H)₂][PF₆] revealed that the equatorial and axial Ir-N bond lengths differed by less than 0.01 Å relative to [Ir(H)₂(PCy₃)(py)₃][BF₄]. However, the former was able to enhance the ¹H NMR signal deriving from the *meta*-protons of pyridine by 266-fold after polarisation transfer in earth's magnetic field, whereas the latter resulted in only an 18-fold increase. Thus, the inclusion of a NHC ligand instead of a phosphine ligand substantially improved the polarisation transfer.

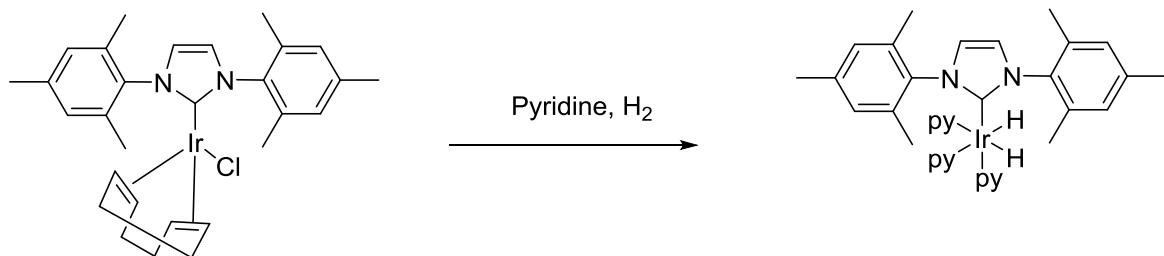


Figure 3. Reaction scheme showing the formation of [Ir(IMes)(py)₃(H)₂]⁺ from [Ir(IMes)(COD)Cl] in the presence of pyridine (py) and H₂.

As SABRE involves the exchange of substrate and *parahydrogen* derived ligands at a metal centre, exchange rates can be measured to give insight into the process. Furthermore, when measured at different temperatures, Eyring plots can be constructed to provide information on the energetics of the system. Exchange rates have been measured using EXSY for [Ir(IMes)(py)₃(H)₂]⁺ and they revealed that at 300 K the dissociation rates of pyridine and hydride are 11.7 s⁻¹ and 9 s⁻¹ respectively [27]. For [Ir(H)₂(PCy₃)(py)₃][BF₄], the analogous pyridine exchange rate is only 1.9 s⁻¹ [25]. These data again highlight that the incorporation of IMes instead of PCy₃ results in the better catalyst. Increasing the pyridine concentration led to a decrease in the observed hydride dissociation rate, which is suggestive that hydride loss occurs after pyridine dissociation.

Optimisation of the SABRE catalyst

Following these early observations, it was clear that the composition of the catalyst was fundamental in terms of the polarisation transfer efficiency. Thus, in an effort to gain insight into optimising the identity of the NHC of the pre-catalyst, $[\text{Ir}(\text{NHC})(\text{COD})\text{Cl}]$, a series of complexes were synthesised [28]. A total of nine complexes were synthesised, of which one was $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$. One complex, $[\text{Ir}(\text{IAd})(\text{COD})\text{Cl}]$, could not be screened due to poor solubility in d_4 -methanol, but the remainder were used to polarise pyridine via SABRE. Plots of signal enhancement against either the modified Tolman parameter, buried volume or exchange rate of pyridine all indicated that no trend was apparent. In all of these plots, $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ was the outlier and notably use of this complex produced the best ^1H NMR signal enhancement of pyridine out of the complexes studied. However, it was shown that increasing the buried volume led to an increase in exchange rate, which is consistent with a dissociative mechanism.

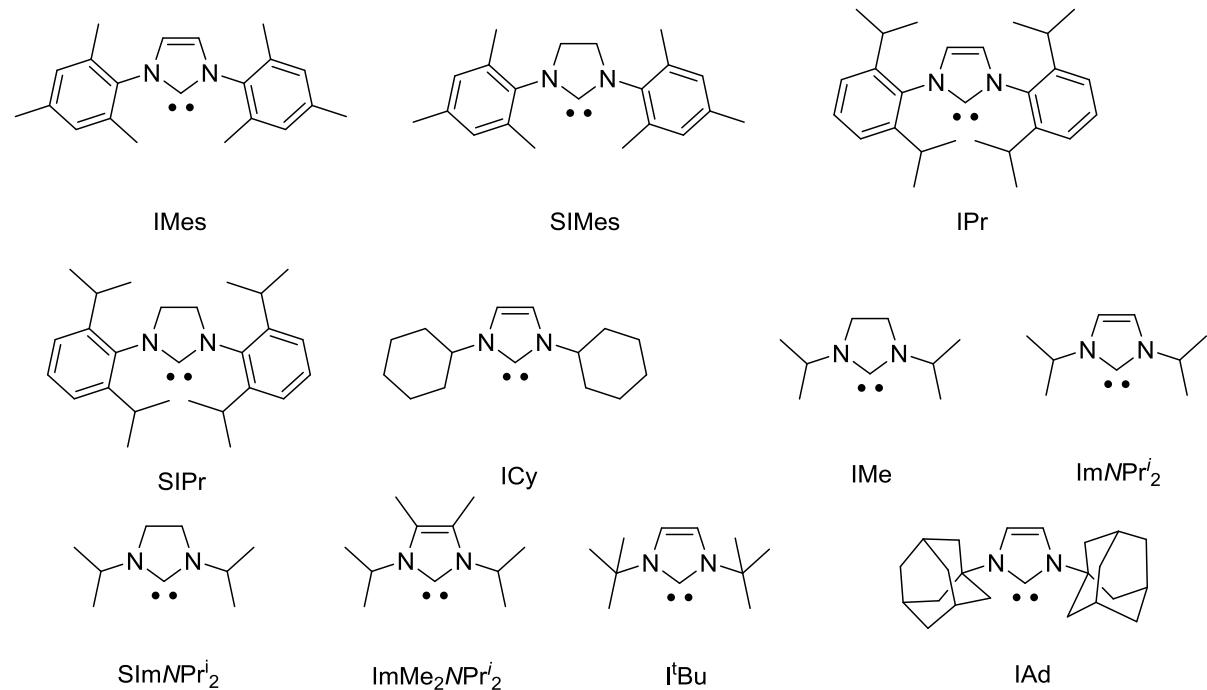


Figure 4. Chemical structures of carbenes evaluated for the production of SABRE catalysts.

A further work on the identity of the carbene in the SABRE process has been published [29]. The pre-catalyst $[\text{Ir}(\text{NHC})(\text{COD})\text{Cl}]$ was again investigated; the NHCs used were IMes, SIMes, IPr, SIPr, ICy, IMe and $\text{ImMe}_2\text{NPr}'_2$. A number of observations were made using pyridine as the test substrate, which gave complexes of the type $[\text{Ir}(\text{NHC})(\text{py})_3(\text{H})_2]^+$ in the presence of hydrogen. Firstly, ligand exchange studies revealed that the loss of pyridine from $[\text{Ir}(\text{NHC})(\text{H})_2(\text{py})_3]^+$ was dissociative and controlled by the steric and electronic parameters of the NHC. When the NHC was bulky, this caused the Ir-py interaction to weaken, therefore promoting pyridine dissociation. Based on entropy values, the trend was found to be $\text{SiPr} > \text{IPr} > \text{SIMes} > \text{IMes} > \text{ImMe}_2\text{NPr}'_2 > \text{ICy} > \text{IMe}$. Due to this dissociative process, $[\text{Ir}(\text{NHC})(\text{H})_2(\text{py})_2(\text{methanol})]^+$ was detected, which had been postulated by prior DFT (Density Functional Theory) studies [27]. While it is SABRE active, the concentration of this species is low and therefore has a minimal role in polarising pyridine. The rate of H_2 loss was also examined and this time the trend was shown to be $\text{SiPr} > \text{SIMes} > \text{IPr} > \text{IMes} > \text{ImMe}_2\text{NPr}'_2 > \text{ICy} > \text{IMe}$. Finally, the different carbene containing catalysts were ranked in terms of delivering the biggest enhancement of pyridine at 300 K. The trend for this is $\text{SiPr} < \text{ICy} < \text{IMe} < \text{ImMe}_2\text{NPr}'_2 < \text{IPr} < \text{SIMes} < \text{IMes}$. It is noteworthy that IMes, which produces the largest enhancement of *ca.* 6000-fold for

pyridine (when a ratio of 1 : 4 [Ir(IMes)(COD)Cl] : py was employed), resides in the middle of the trends examined and thus suggests that an optimum value for the dissociation rate of pyridine and H₂ exists. The SABRE catalyst can, therefore, be further optimised and by doing so, deliver far greater signal enhancements.

The π-accepting ability parameter (PAAP) appears to be one of the better parameters to analyse the electronic properties of NHC ligands [30]. The chemical shift of ⁷⁷Se in selonourea complexes of various NHCs (SImNPr, ImNPr, ICy, SIMes, IMes, SIPr, IPr and IPent) was plotted against exchange rate of pyridine (Figure 5). This revealed a strong correlation between the two parameters. However, there were anomalies. Firstly, ^tBu was not utilised in the comparison as its extremely high ⁷⁷Se chemical shift (δ 183) has been already considered to be anomalous [31]. Furthermore, SImNPr did not fit along the correlation curve; despite having an exchange that is three times larger than its unsaturated analogue, ImNPr. The low absolute exchange rate of SImNPr was attributed to the aliphatic nature of the isopropyl group and that it has a relatively small buried volume (26.4%). The effect of steric bulk was exemplified further by comparing IPr to IPent. The resulting selonourea complexes have comparative chemical shift values for ⁷⁷Se (δ 90 for IPr and δ 101 for IPent). The buried volume of IPent is, however, approximately 5% larger than IPr. Based on the general trend (derived from second order polynomial expression, see Figure 5), an exchange rate for pyridine for IPent was calculated to be 41.6 s⁻¹. As the bound pyridine signal was only observable at temperatures below 263 K, a much higher exchange rate was expected. Experimentally, the pyridine exchange rate for IPent was not determined. Thus, although the PAAP is useful in terms of ascertaining whether an NHC is applicable for SABRE, the steric bulk and other parameters should be considered also. It is important to note that IMes, the NHC that produces the best signal enhancement, is towards the bottom of the curve.

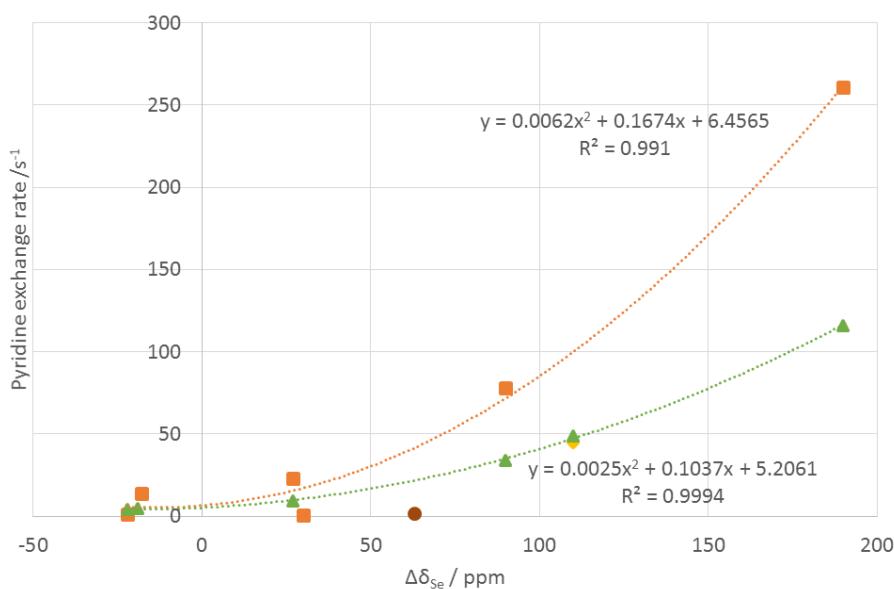


Figure 5. Plot of pyridine exchange rate against $\Delta\delta_{\text{Se}}$ using data reported by Duckett and co-workers and Tessari and co-workers. ■ Data reported by Duckett and co-workers at 300 K. ◆ SIMes data point from Duckett and co-workers (outlier). ▲ Data reported by Tessari and co-workers at 298 K. ● SImNPr₂ data point from Tessari and co-workers (outlier)

Exchange data, in conjunction with DFT, have been used to rationalise the exchange processes occurring at the iridium centre of [Ir(IMes)(py)₃(H)]⁺ and hence the mechanism of polarisation transfer. Optimisation of the [Ir(IMes)(py)₃(H)]⁺ at the BP86/LANL2DZ level resulted in geometrical parameters very close to those determined from experimentally from X-ray crystallographic data

[27]. Loss of pyridine *trans* to hydride or *trans* to IMes from $[\text{Ir}(\text{IMes})(\text{py})_3(\text{H})_2]^+$ resulted in the same product after reorganization. Experimentally, EXSY data showed that only the pyridine *trans* to hydride dissociates on the NMR timescale. The computed gas-phase free-energy barrier for loss of pyridine *trans* to hydride is 48 kJ mol⁻¹; this compares well with the experimentally determined value of 64 ± 2 kJ mol⁻¹. The positive ΔS value of 97 ± 9 J K⁻¹ mol⁻¹ for this process determined experimentally from the Eyring plot also indicated a dissociative mechanism. Three different routes of H₂ exchange were investigated. Loss of dihydride from the intermediate, $[\text{Ir}(\text{IMes})(\text{py})_2(\text{H})_2]\text{Cl}$, to give the 14-electron species $[\text{Ir}(\text{IMes})(\text{py})_2]\text{Cl}$, was shown to be thermodynamically unfavourable, as at least 117 kJ mol⁻¹ of free energy is required. The second route involves solvent participation to yield an 18-electron intermediate, $[\text{Ir}(\text{IMes})(\text{py})_2(\text{H})_2(\text{MeOH})]\text{Cl}$. Whilst methanol binding requires only 13 kJ mol⁻¹ of free energy, subsequent elimination of dihydride requires an additional 83 kJ mol⁻¹. Again, the experimental data did not support these findings. The third route, and most likely based on experimental evidence, involves the formation of $[\text{Ir}(\text{IMes})(\text{py})_3(\text{H})_2(\eta^2\text{-H}_2)]^+$. This involves the incoming H₂ ligand to bind in the faster relaxing dihydrogen form. The combined gas-phase free-energy barrier to form this product from $[\text{Ir}(\text{IMes})(\text{py})_3(\text{H})_2]^+$ is 44 kJ mol⁻¹. However, this would require the lifetime of $[\text{Ir}(\text{IMes})(\text{py})_3(\text{H})_2(\eta^2\text{-H}_2)]^+$ to be short enough to avoid the conversion of *parahydrogen* to *ortho*hydrogen.

To gain further insight into the polarisation transfer process, and hence improve the catalytic system for future experiments, DFT was again used to provide a computational basis for the presence of various complexes produced when $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ was exposed to methyltriazole (mtz), pyridine and H₂ [32]. The complexes possibly formed by this process are shown in Figure 6 along with their relative energies. The relative energies could not be directly compared to previous work, although the optimized geometry of $[\text{Ir}(\text{IMes})(\text{py})_3(\text{H})_2]^+$ does show similar features. $[\text{Ir}(\text{IMes})(\text{py})_2(\text{mtz})(\text{H})_2]^+$ has the highest relative energy of the complexes investigated. This is consistent with NMR data, as this species is not observable using this technique. The relative abundances at different ratios of pyridine : mtz were computed using DFT and compared with NMR data. This revealed that $[\text{Ir}(\text{IMes})(\text{py})_3(\text{H})_2]^+$ was only observed at very low concentrations of mtz (below 1 : 0.6 pyridine : mtz). DFT calculations indicated that the presence of two structural isomers of $[\text{Ir}(\text{IMes})(\text{py})(\text{mtz})_2(\text{H})_2]^+$, in which both mtz ligands are *trans* to hydride or only one mtz ligand is *trans* to hydride and the other *trans* to IMes, were underestimated and overestimated respectively relative to NMR experimental data. Similarly, $[\text{Ir}(\text{IMes})(\text{mtz})_3(\text{H})_2]^+$ was underestimated. The relative concentrations of $[\text{Ir}(\text{IMes})(\text{py})_2(\text{mtz})(\text{H})_2]^+$ (one pyridine *trans* to IMes, the other *trans* to hydride) and $[\text{Ir}(\text{IMes})(\text{py})(\text{mtz})_2(\text{H})_2]^+$ (one mtz *trans* to IMes, the other *trans* to hydride) were shown to be in good agreement, albeit the former more than the latter, with the experimentally determined ¹H NMR data. The coordination bonds in these complexes were subjected to energy decomposition analysis (EDA). The EDA performed for the ligand *trans* to NHC and both *cis*-ligands revealed that the *trans* ligand was hardly affected by the nature of the *cis* ligands, and *vice versa*. Furthermore, regardless of the type of ligand, *trans* ligands produced a more stabilizing interaction than *cis* ligands. This was rationalised to be due to more favourable electrostatic and orbital interaction energy terms in *trans* ligands compared to those of *cis* ligands.

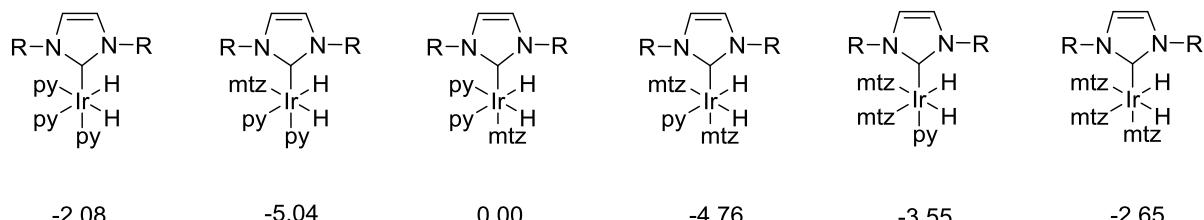


Figure 6. Chemical structures (py = pyridine, mtz = methyltriazole and R = 1,3,6-trimethylphenyl) and their relative energies (kcal mol⁻¹) calculated by DFT [32].

A simple analytical model for SABRE has been presented that provides rationale for deciding which system parameters to optimise to achieve higher signal enhancements [33]. Parameters such as *J*-couplings, relaxation rates and reaction rate constants were considered. An expression was derived for determining the enhancement provided by SABRE. The relaxation rates were highlighted as being the key parameter to be optimised and not the dissociation constants. Furthermore, an optimal value for the dissociation rate constant for the substrate was derived to be 4.5 s⁻¹ (which equalled the angular frequency of a level anti-crossing (LAC)), ω_{LAC} , which is lower than that of pyridine in $[\text{Ir}(\text{IMes})(\text{py})_3(\text{H})_2]^+$ (11.7 s⁻¹) [27]. A theoretical approach based on LACs had previously been described for pyridine polarised by Crabtree's catalyst which accounted for the SABRE effects and their magnetic-field dependence observed experimentally [34]. Thus, the expression could be used to predict the structure of future catalysts for SABRE. Recommendations for the design of an optimal chemical system were proposed and these were : i) the substrate-to-catalyst ratio should be as low as possible provided that $[\text{Ir}(\text{NHC})(\text{H})_2(\text{sub})_3]^+$ remains the main catalyst in solution; ii) for low-field SABRE, the substrate dissociation rate constant should be of the same order as the spin mixing frequency, whereas for high-field sample, the dissociation rate should be as high as possible; iii) the catalytic species should have minimal effect on the relaxation rate; iv) the *parahydrogen* concentration in solution should be as high as possible.

SABRE catalysts utilising rhodium and cobalt

To date only one example of a rhodium based catalyst being used in conjunction with SABRE has been reported. This is despite numerous reports of its use in PHIP-based studies [11-13, 35]. Zhivonitko *et al.* have demonstrated that $[\text{Ir}(\text{PPh}_3)_3(\text{H})_2\text{Cl}]$ can be generated easily from $[\text{Ir}_2\text{Cl}_2(\text{COD})_2]$ and PPh_3 in the presence of *parahydrogen* to produce hyperpolarised ³¹P NMR spectra (see the ³¹P Imaging Applications of SABRE section for more details) [36]. They further report that $[\text{Ir}(\text{COD})(\text{PCy}_3)(\text{py})]\text{PF}_6$ and $[\text{Rh}(\text{COD})(\text{DPPB})]\text{BF}_4$ (DPPB = 1,4-*bis*(diphenylphosphino)butane) also demonstrate polarisation transfer *via* SABRE to ³¹P. However, the authors note that the signals were hard to assign to a particular metal complex most likely due to the complex equilibria present and the low concentrations of polarised species. To date, this is the only report of rhodium complexes being used for SABRE and, therefore, represents a major area for future development.

Recent literature has also demonstrated the use of cobalt to facilitate polarisation transfer in SABRE. The majority of extant SABRE literature focuses on the use of a range of iridium centred polarisation transfer catalysts, however recent publications have demonstrated cobalt catalysts can also facilitate the SABRE effect. The first reported instance made use of $(\text{Mes}^{\text{CCC}})\text{Co}(\text{N}_2)(\text{PPh}_3)$, a trigonal dipyramidal complex supported by an electron rich monoanionic pincer ligand Mes^{CCC} (*bis*(mesityl-benzimidazol-2-ylidene)phenyl) [37]. This *bis*-carbene ligand is structurally similar to the commonly used electron donating NHC IMes, whose advantages over previously used phosphines has been discussed [27]. Mes^{CCC} predominantly differs from IMes in terms of the buried volume, one of the key parameters for the optimisation of the SABRE process, which could restrict bulky substrates from making use of this system. Studies of ligand buried volume in SABRE catalysis have

demonstrated that ligand steric properties can impact catalytic activity as well as determining an optimal buried volume which when exceeded results in a loss of signal enhancement [28]. However, the literature of SABRE ligand optimisation focuses almost exclusively on monodentate ligands as opposed to the tridentate system described by Tokmic *et al.* [37]. It should be noted, though, that in the proposed octahedral coordination geometry the presence of the tridentate NHC and two hydrides will likely restrict the exchange rate of the substrate on the catalyst thus reducing the build-up rate of hyperpolarised bolus.

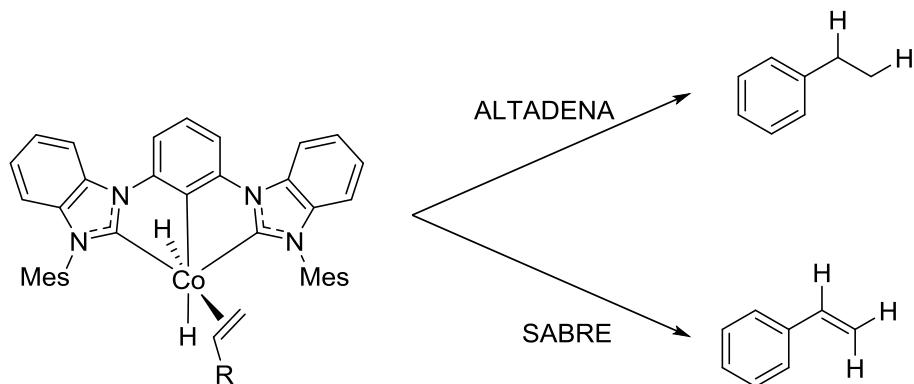


Figure 7. Styrene products identified when the indicated cobalt catalyst, in conjunction with *parahydrogen*, was utilised. R = Ph and Mes = 1,3,6-trimethylphenyl

The cobalt catalytic system was intended to facilitate hydrogenation of olefins, which would be monitored making use of PHIP, however, SABRE was observed in a joint manner along with the expected hydrogenation under 4 atmospheres of 99.99% *parahydrogen* in a standard Young's NMR tube. OPSY (only *parahydrogen* spectroscopy) [38, 39] showed that this effect occurs for a range of substrates (including styrene, 4-vinylcyclohexene and both geometric isomers of 4-octene) [40] all demonstrating simultaneous PHIP and SABRE enhancements (see Figure 7 for an example). All the substrates used demonstrate only the terminal alkene underwent PHIP whereas SABRE was observed at all alkene sites (and in neighbouring proton environments in olefins) except for the *ortho*-protons in styrene. Furthermore, the reversible exchange of hydrogen without hydrogenation at the alkene sites was demonstrated through the spectroscopic observation of HD gas when the reaction was carried out under deuterium (4 atm).

SABRE catalysts with bi- and tri-dentate ligands

Iridium catalysts have been prepared with bi- and tri-dentate ligands, which resemble the cobalt catalysts shown in the previous section. A SABRE catalyst, which utilises a bidentate carbene ligand, has been synthesised by Ruddlesden *et al.* and the solvent capability of SABRE investigated [41]. In d_2 -DCM, an enhancement of 141-fold was observed for the *ortho*-proton resonances. This increases to 346-fold after a period of 24 days due to the slow formation of the active form of the catalyst. The complex shown in Figure 8 is the only inorganic product detected and no deuterium exchange with the solvent is evident. Similar observations were made in benzene and THF, with THF yielding a total enhancement of *ca.* 1300-fold after 24 days. When the solvent was changed to methanol or ethanol, different behaviour was observed. After 24 days of exposure, significant deuterium incorporation was observed for the *ortho* signal of pyridine. In the case of methanol, 50% deuteration had occurred. The same process occurs for ethanol but is slower. Thus, when SABRE was conducted in methanol, the *ortho* signal is enhanced by 185-fold, but after 24 days, the signal decreased to 29-fold. Although for ethanol, the analogous enhancements are 86- and 282-fold respectively, these are still much smaller than those observed for the non-protic solvents.

Table 2. ^1H NMR signal enhancements for the three ^1H sites of pyridine (0.15 M) following polarisation transfer at 65 G and using the [Ir(IMes)(COD)Cl] template (15 mol%) after 48 hours and 24 days

	^1H NMR enhancement following SABRE for the indicated pyridine sites					
	48 h			24 days		
Deuterated solvent	<i>ortho</i> / fold	<i>meta</i> / fold	<i>para</i> / fold	<i>ortho</i> / fold	<i>meta</i> / fold	<i>para</i> / fold
Benzene	121	93	56	561	469	253
THF	98	61	49	600	442	265
DCM	141	105	65	346	212	159
Ethanol	86	35	36	282	145	142
Methanol	185	130	82	29	24	18

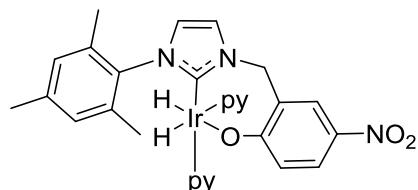


Figure 8. The inorganic product detected in d_2 -DCM by Ruddlesden *et al.* that is responsible for polarisation transfer [41].

Modification of the catalyst structure as shown in Figure 8 was modified so that it did not possess a nitro group on the phenolate ring [42]. The Ir-O bond is retained in d_6 -benzene and d_2 -DCM in the presence of pyridine and hydrogen. However, in methanol, the Ir-O bond breaks and subsequently, the phenolate dissociates to form the zwitterionic complex. A single hydride at δ -22.18 is observed for this complex. Pyridine, nicotinaldehyde and nicotine were all polarised in the three solvents described. Enhancements were consistently better in DCM and benzene with an enhancement of 1850-fold over the five proton sites of pyridine being obtained. In DCM and benzene, only one substrate ligand is ligated *trans* to hydride, whereas in methanol, there are two because the phenolate is no longer ligated to the metal centre. Thus, the available polarisation becomes shared over a greater number of spins and, therefore, the signal enhancement is smaller. Both this report and the one detailed in ref. [41] highlight the importance of solvent selection in terms of maximising the observable enhancement due to polarisation transfer.

A SABRE catalyst that utilises a tri-dentate ligand, a PNP pincer ligand, has also been synthesised and the polarisation transfer to pyridine investigated [43]. After polarisation transfer in a magnetic field of 65 G, the *ortho*-proton signal of pyridine was enhanced by 12-fold. Warming the solution from 295 K to 310 K increases the signal intensity further (by 8%). The pyridine exchange rates at 298 and 310 K are $18 \pm 4 \text{ s}^{-1}$ and $66 \pm 18 \text{ s}^{-1}$; the former is more suited to polarisation transfer at 0 G whereas the latter to 65 G. In either case, the efficiency of the transfer is not as good as [Ir(IMes)(COD)Cl] [27, 29], for example. Deuterating the PNP ligand, as shown in Figure 9, resulted in a 35% increase in efficiency of polarisation transfer to pyridine at 65 G; the *ortho*-protons were enhanced 16-fold respectively using the deuterated template.

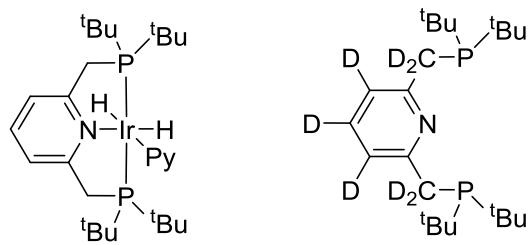


Figure 9. Chemical structures of the PNP pincer based catalyst and the deuterated pincer ligand synthesised by Holmes *et al* [43].

Exploring the field dependency of SABRE

The magnetic field in which polarisation is transferred is key to optimising the coherent spin mixing which is the dominant polarisation transfer mechanism. Maximising coherent spin mixing thus entails larger enhancements. The seemingly similar magnetic field dependence of SABRE has been examined using a range of substrates [44]. The substrates pyridine, pyrazole, 1-methylpyrazole, 3-methylpyrazole, 3-phenylpyrazole, 3-(3-bromophenyl)pyrazole, indazole, 3,5-dimethylpyrazole and 3,5-diphenylpyrazole were all examined in this work and were polarised using the $[\text{Ir}(\text{PCy}_3)(\text{COD})(\text{py})]$ template. All of the substrates analysed, aside from 3,5-dimethylpyrazole, indazole and 3,5-diphenylpyrazole, possessed similar magnetic field dependences in terms of maximal enhancements. Maxima were observed at 3 – 4 mT and 12 – 14 mT with a 180° phase difference apparent. The magnetic field dependence of indazole was not determined as the signal enhancement was too low. 3,5-diphenylpyrazole was not enhanced under the conditions employed. As the pyridine ligand of the catalyst was also not polarised, this observation is suggestive of an inactive iridium complex being formed. The phenyl rings of 3-phenylpyrazole and 3-(3-bromophenyl)pyrazole both showed no signal enhancement whereas the pyrazole ring protons did so. 3,5-dimethylpyrazole possessed a different magnetic field dependence. It is noteworthy that only a ${}^5J_{\text{HH}}$ coupling exists between a *parahydrogen* derived hydride and its aromatic proton. This contrasts to the other substrates as they possess ${}^4J_{\text{HH}}$ couplings. The aromatic proton of 3,5-dimethylpyrazole was only enhanced following polarisation transfer at magnetic fields above 5 mT whereas the methyl groups showed no enhancement over the field range 0 – 25 mT. This work also noted that the hydroxyl ${}^1\text{H}$ NMR signal of CD_3OH was slightly polarised when pyrazole was investigated.

In a further study, Moreno *et al.* probed the field dependency of the enhancement observed for the hydroxyl ${}^1\text{H}$ NMR signal of CD_3OH [45]. In the absence of HCl, the hydroxyl signal showed negligible polarisation when pyridine was polarised in the presence of the pre-catalyst $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ and *parahydrogen*. However, the addition of aqueous HCl to a concentration of 5 mM, led to a 40-fold enhancement at 25 mT, which mirrored the magnetic field in which optimal enhancement was observed for pyridine. Exchanging the substrate for pyrazole or pyridazine led to the same observation, albeit both enhancement values were smaller. The employment of imidazole instead of pyridine resulted in no enhancement of the hydroxyl ${}^1\text{H}$ NMR signal of CD_3OH . When d_5 -pyridine was used instead of protio-pyridine, the enhancement of the solvent dropped by a factor of six, suggesting the polarisation transfer to the solvent occurs through chemical exchange or by direct coordination of the solvent to the catalyst to propagate polarisation. Cross-relaxation from the hydride protons to bound solvent molecules may also occur at high-field.

The reasons behind the observed field dependency similarity of SABRE have been explored theoretically and compared to experimental results [46]. At low fields, the dominant polarisation transfer mechanism for SABRE is coherent spin mixing, and this mixing is the most efficient in LAC regions. By comparing the use of Crabtree's catalyst and $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ to polarise pyridine, it was

possible to explain the reason why two maxima are observed in the field dependence plot over the region 0 – 50 mT for the former, whereas only one maxima was observed for the latter. Spin-spin interactions in Crabtree's catalyst involving the ^{31}P spins caused the LAC to split, thus doubling the number of maxima observed. These observations demonstrate the significant role of spin- $\frac{1}{2}$ heteronuclei in the formation of SABRE at low magnetic fields. The experimental and theoretical curves produced were in very good agreement and the model further suggested that only the *ortho*-protons of pyridine acquire polarisation directly from *parahydrogen* derived hydrides; the other protons acquire polarisation indirectly from the *ortho*-proton *via* their *J*-coupling.

The design, and employment of a fast field-cycling device has enabled the field dependency of the SABRE enhancement of 1,2,3-triazole to be measured and compared to theoretical values [47]. The pre-catalyst [Ir(IMes)(COD)Cl] was employed. The field cycling device enabled the transfer of the hyperpolarised solution from the polarisation transfer field to the measurement field in 0.5 s, which is quicker than the conventional transfer method of inserting the sample manually (generally in the order of seconds). Resultantly, the field dependency of 1,2,3-triazole showed a single maxima at 6 mT with a corresponding enhancement value of 600-fold. The experimental and theoretical data showed very good agreement, and again showed that a single LAC region existed over the field range explored (0 – 15 mT). A single LAC was also observed for the hydride ligands of the polarisation catalyst, and the maxima was again observed at 6 mT, although unlike 1,2,3-triazole, this maxima was positive rather than negative.

Automated delivery of SABRE hyperpolarised samples using a Polarizer

The nature of the polarisation transfer in SABRE, namely shaking the tube in a fringe field of a superconducting magnet, before transfer into the measurement field, has led some to groups to evaluate other strategies to transfer the hyperpolarised solution into the magnet. Duckett and co-workers have described how a reaction cell has been interfaced with an NMR spectrometer (equipped with a flow NMR probe that has a head flow volume of 200 μL) to measure both the sample hyperpolarisation level and the nuclear spin states [27]. The reaction cell is housed within a magnetic coil that is capable of generating a magnetic field up to 150 G in the vertical direction only. The setup enables a predefined magnetic field to be established prior to bubbling *parahydrogen* through a solution for a set period of time (10 s was used at a pressure of 2 bar). Using this setup, ^1H NMR spectra were obtained for hyperpolarised pyridine over the range of 0 (magnetic coil turned off) – 150 G. The *ortho*- and *para*- peaks mirror one another in terms of the direction (all emission) and relative size of the signals produced. A maximum was observed at 65 – 70 G. The *meta*-signal, also had a maximum enhancement following polarisation transfer at 65 – 70 G, however, the signal direction flipped to being purely absorption at 15 G from emission.

The reaction cell linked to a flow-probe (often referred to as the Polarizer) can be fully automated when interfaced with the NMR hardware correctly and the appropriate pulse sequences used [48]. This circumvents experimental error associated with initiating polarisation transfer by shaking the sample tube and also the transfer of the sample to the measurement field. This setup has been used to probe the magnetic states of nicotinamide, and to optimise the conditions under which polarisation can be transferred most effectively. Increasing the *parahydrogen* pressure from 0.5 to 5 bar (limit of the equipment) showed an increase in the signal intensity from *ca.* 150-fold to *ca.* 850-fold for the four proton environments of nicotinamide. These measurements were conducted using a *parahydrogen* purge time (the time that *parahydrogen* is flowed through the reaction chamber) of 6 s. In a separate experiment, increasing the *parahydrogen* purge time from 1 to 30 s showed a

general increase with evidence for plateauing between 20 and 30 s. The concentration of the polarisation transfer catalyst, $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$, was found to affect the build-up of polarisation in nicotinamide and the resulting enhancement. This is because the catalyst not only facilitates polarisation transfer, but also accelerates T_1 . Both of these factors are concentration dependent. Using the Polarizer, the production of higher order spin terms *via* SABRE was probed using OPSY [38, 39]. The hard pulses of the OPSY sequence were exchanged for selective pulses so that individual resonances could be excited. Thus, the longitudinal four-spin order term, the four longitudinal three-spin order terms and the six longitudinal two-spin order terms were all detected and their relative signal enhancements reported. The relative amplitudes decrease with the increasing number of spins involved in the spin order term.

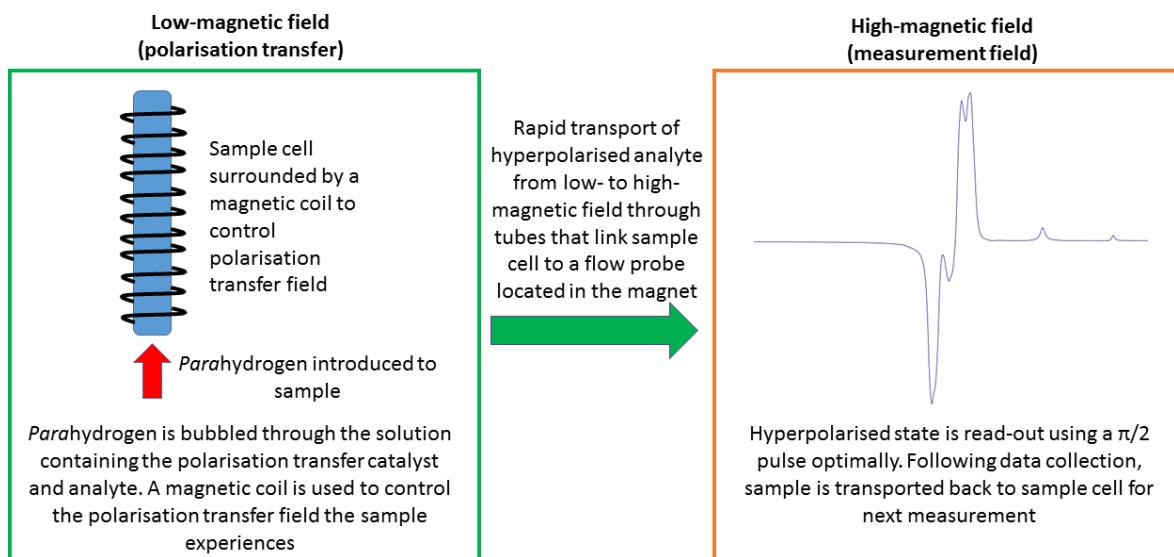


Figure 10. Overview of the Polarizer

High-field SABRE

Normally, polarisation transfer from *parahydrogen* derived hydrides to nuclei in a desired analyte is transferred at low magnetic field (typically less than 150 G). Polarisation within the field of the spectrometer (e.g. 9.4 T) was originally thought not to result in SABRE. However, Chekmenev and co-workers have shown that it is possible to generate SABRE-derived signals in a high-field instrumentation using the $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ template [49]. Compared to performing the polarisation transfer at low-field, only the *ortho*-signals of pyridine (4-fold enhancement) were significantly enhanced following 2 minutes of *parahydrogen* being bubbled through the solution. The rate of signal build-up, T_b , was found to be 21 ± 2 s which mirrors that of the T_1 (19 ± 1 s). Nuclear spin cross-relaxation was suggested as the mechanism for polarisation transfer. This was later confirmed by Pravdivstev *et al.* through a theoretical and experimental approach [50]. The analysis of d_5 -pyridine also highlighted that H/D exchange was occurring due to the formation of d_4 -pyridine over a number of hours. H/D exchange is unlikely to contribute significantly to the hyperpolarisation process.

High-field SABRE has been further investigated using RF saturation studies [51]. Application of a frequency-selective saturation RF pulse at the hyperpolarised hydride resonance frequency during

polarisation transfer resulted in diminished signal intensities for the *ortho*-hydrogen peak and the hydride peak even when using a soft ($B \sim 7.5$ Hz) RF saturation. In addition, the use of a relatively strong RF saturation ($B \sim 2.4 \times 10^2$ Hz) resulted in significant reduction of the high-field SABRE enhancement. Similar effects were seen by applying the RF saturation pulse to the *ortho*hydrogen signal, whereas shifting the RF saturation to the *ortho* ^1H NMR signal of pyridine resulted in suppression of the *ortho* free pyridine ^1H NMR signal and that of the exchangeable pyridine signal bound to the catalyst. These results reinforce the proposed nuclear dipolar cross-relaxation mechanism for polarisation transfer.

It has subsequently been demonstrated that high-field SABRE can be much improved by using a spin-lock [52, 53]. Following bubbling a sample containing $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ and pyridine with *parahydrogen* gas, a RF-field with amplitude v and frequency v_{rf} was turned on for τ_{rf} (which is smaller than the relaxation time of the protons in free pyridine and comparable to the relaxation time of the pyridine protons in the active complex). Enhancements of *ca.* 360 were obtained for the single proton of 3,4,5,6-*d*₄-pyridine which are larger than those of high-field SABRE without RF excitation [49]. Pyridine, 3-methyl-1H-pyrazole, 4,4'-bipyridine and 2,2'-bipyridine were also polarised using this technique. The method exploits the mixing of spin states at Level Anti-Crossings (LACs) as identified from the doubly tilted frame (DTF) schematic representations; fulfilling the LAC conditions by utilising the appropriate RF-field results in the observation of polarisation transfer. The phase of the spectrum can be controlled by using the appropriate v_{rf} to select the corresponding LAC.

High-field SABRE has been combined with a continuous hyperpolarisation approach to collect 2D ^1H NMR spectra of complex mixtures in the nanomolar range [54]. The pulse sequence used is detailed in Figure 11 and utilised a 1 s *parahydrogen* bubbling time. This approach specifically focused on the bound substrate and hydride signals generated and as such, did not require stringent conditions such as the employment of selective radiofrequency irradiation to collect the data. A ^1H - ^1H COSY NMR spectrum was acquired in 23 minutes of a mixture of thirteen substrates at concentrations between 250 nM and 2 μM . Cross peaks were generated between the substrate proton nuclei and those of the *parahydrogen* derived hydrides which are typically $^4J_{\text{HH}}$. Compared to the corresponding 1D ^1H NMR spectrum of the aromatic region (which took 9.5 hours of measurement time to obtain a sufficient signal-to-noise ratio for analysis), all the hydride signals are resolved in the COSY spectrum and thus this approach circumvents extensive signal crowding. Linear plots of the 2D signal intensities were utilised in standard addition curves which enabled the concentrations of substrates, such as (-)-cotinine and methyl nicotinate, to be determined in good agreement with the nominal values.

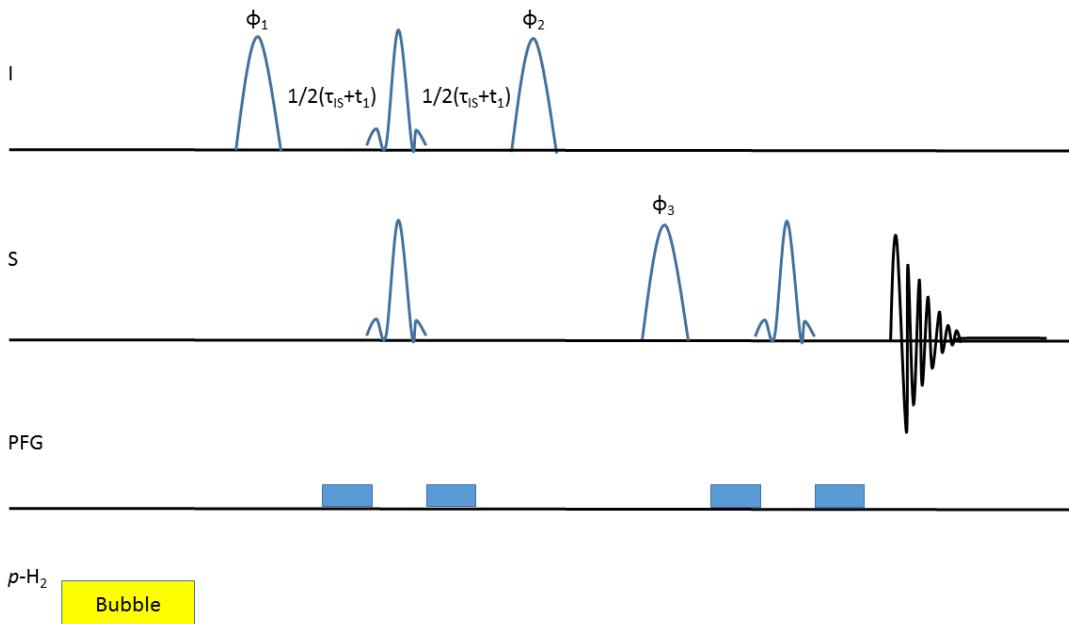


Figure 11. Pulse scheme for continuous hyperpolarisation at high-field developed by Tessari and co-workers [54]. The different ϕ values are for the phase cycling implemented. τ_s is $\frac{1}{2} J_{AX}$, where $J_{AX} = 8$ Hz (hydride-hydride coupling)

This approach has been extended to provide NMR-based chemosensing of natural extracts, namely those contained within ground coffee [55]. High-field *parahydrogen* enhanced 2D correlation spectra linked the hydride signals with substrates bound *trans* to them in the $[\text{Ir}(\text{IMes})(\text{H})_2(\text{sub})(\text{mtz})_2]^+$ complexes that formed (mtz used as a co-substrate). This again simplified spectra to overcome spectral crowding. The chemosensory response is effectively the generation of two hydride signals that results when a substrate species associates with the complex, which are unique to that complex. Furthermore, as the hydride signals are derived from *parahydrogen*, substantial signal-to-noise results. Using this approach, in conjunction with standard addition curves, substrates such as 3-hydroxypyridine, pyridine, pyrazine and 2-methylpyrazine were detected in the micromolar regime.

A further high-field study using 2D NMR has been reported that focuses on heteronuclear data collection [56]. A pulse sequence was developed, SABRE-HETCOR, that makes use of the SABRE-INEPT sequence [25]. Unlike PH-INEPT and PH-INEPT+, which transfer polarisation only to the bound substrate, SABRE-INEPT, enables efficient transfer to the free form of the substrate. This was exemplified for ^{15}N -pyridine and $^{15}\text{N}_2$ -imidazole. Using a theoretical description of the underlying spin dynamics and chemical kinetics of SABRE [57], the ^{15}N polarisation of ^{15}N -pyridine formed via the SABRE-INEPT pulse sequence was probed. Experimentally, the ^{15}N signal of free pyridine was enhanced by 200 fold and the lifetime of the hyperpolarisation calculated to be either 10 s for when [free-pyridine] : [equatorially bound pyridine] is 10 : 1 or 100 s for when [free-pyridine] : [equatorially bound pyridine] is 20 : 1. A $^{15}\text{N}-^1\text{H}$ 2D NMR spectrum obtained using SABRE-HETCOR revealed three cross-peaks for a solution of 2mM ^{15}N -pyridine in the presence of hydrogen and the pre-catalyst $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$; two for equatorially bound pyridine and one for free-pyridine. This spectrum was collected in 64 transients with 8 scans per transient. An advantage of the technique is not having to wait several T_1 periods prior to collecting the next scan; simply bubbling *parahydrogen* through the solution for a few seconds is sufficient to begin a new acquisition.

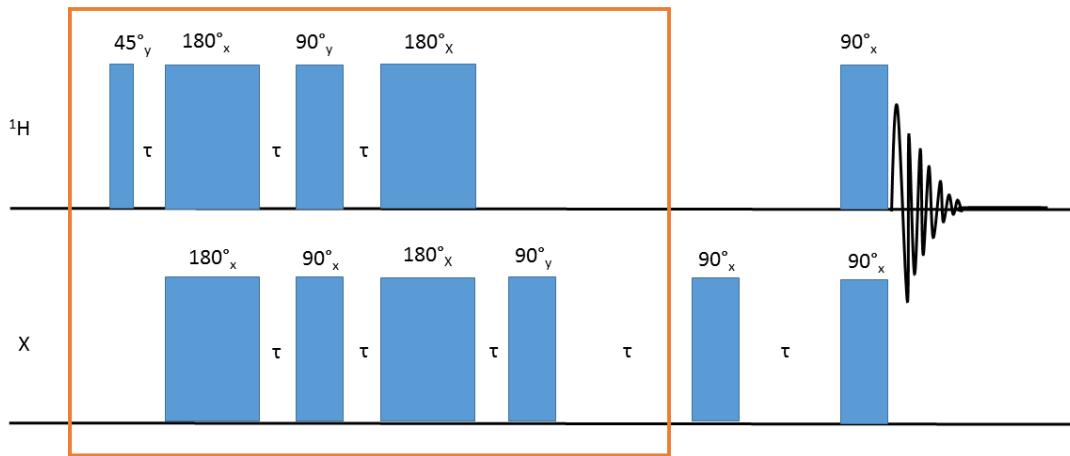


Figure 12. SABRE-HETCOR pulse sequence. The boxed component is the SABRE-INEPT sequence.

The $[\text{Ir}(\text{IMes})(\text{H})_2(\text{mtz})_2(\text{py})]^+$ system, along with $[\text{Ir}(\text{SIMes})(\text{H})_2(\text{mtz})_2(\text{py})]^+$, have been the focus of a study in which EXSY data were collected in a short experimental time using a high-field hyperpolarisation-EXSY pulse sequence [58]. As the pulse sequence relies upon the hyperpolarisation of the hydrides, there is no requirement for excessively long recycle delays. Enhanced substrate proton magnetisation is obtained following a COSY-type coherence transfer from the hyperpolarised hydride ligands. The magnetisation of the bound substrate is *ca.* 100 times larger than thermal magnetisation. As the pulse sequence employs phase cycling only signals originating from the hyperpolarised hydrides are detected. Using the sequence meant that for a whole series of exchange periods, the total experiment time was only ten minutes. In addition, the high signal-to-noise ratio obtained due to hyperpolarisation, meant that relatively dilute solutions (e.g. 250 μM of pyridine) could be studied. Pyridine dissociation rates of 0.19 s^{-1} and 0.66 s^{-1} for $[\text{Ir}(\text{IMes})(\text{H})_2(\text{mtz})_2(\text{py})]^+$ and $[\text{Ir}(\text{SIMes})(\text{H})_2(\text{mtz})_2(\text{py})]^+$ respectively, which are in good agreement with reported estimates [59]. The pulse sequence can be used to measure dissociation rates between 0.1 s^{-1} and 3 s^{-1} , beyond this longitudinal relaxation of the substrate proton magnetisation limits transfer magnetisation to the free substrate. For much faster dissociation rates (i.e. $<0.1 \text{ s}^{-1}$), the efficiency of the COSY-transfer reduces, leading to a fall in experiment sensitivity and signal-to-noise gain offered by hyperpolarisation.

Strong enhancements for pyridine have been observed in ^{15}N NMR spectrum using high-field SABRE [60]. Free pyridine was enhanced by 800-fold whereas the equatorially bound pyridine ligands in $[\text{Ir}(\text{IMes})(\text{py})_3(\text{H})_2]^+$ were enhanced by 12000-fold. Analogous results were obtained when Crabtree's catalyst was employed. An RF field was used to excite the frequency of the hydrides in the ^1H domain whereas the frequency of the ^{15}N RF-field was varied so that the spin system passed through a LAC for coherent spin mixing. The pulse sequence is outlined in Figure 13. As the experiment uses a low RF-excitation bandwidth, polarisation cannot be transferred over a wide range of chemical shifts. However, the method does use low RF powers and so heating of the sample does not occur.

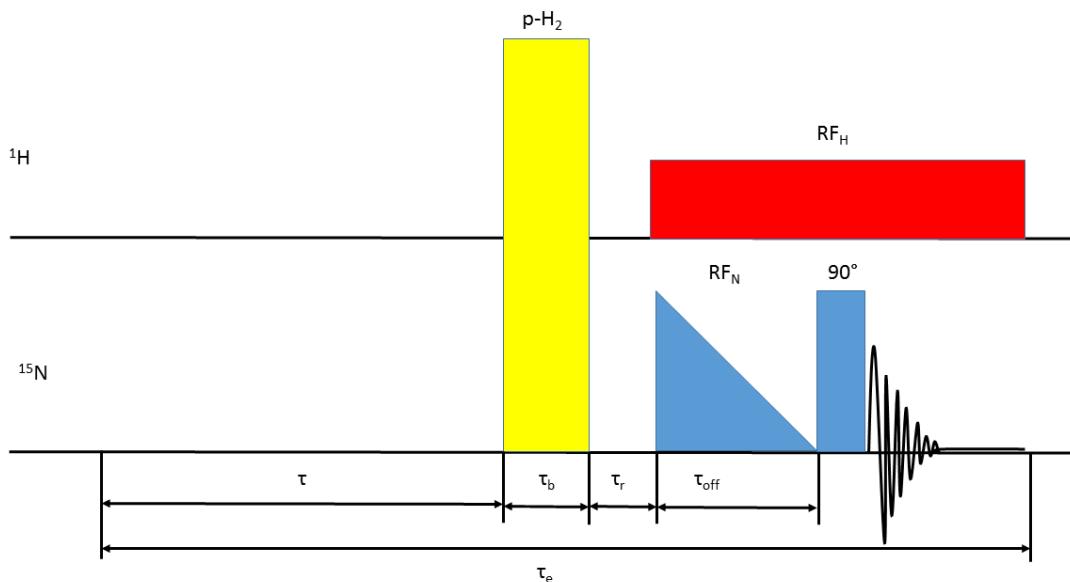


Figure 13. Transfer of polarisation to ^{15}N at high-field developed by Ivanov and co-workers [60]. τ_1 is the time to relax to thermal equilibrium, τ_b is the *parahydrogen* bubbling time, τ_r is the wait time to allow bubbles to escape from solution, τ_{off} is the time taken for the RF_N -field to linearly reduce to zero and τ_e is the total experiment time.

Employment of SABRE to polarise spin- $\frac{1}{2}$ nuclei

Many spin- $\frac{1}{2}$ nuclei have been polarised by SABRE, namely ^1H , ^{13}C , ^{15}N , ^{19}F , ^{31}P , ^{29}Si and ^{119}Sn . This section will look at each nucleus individually and detail the discoveries. The applications for which these nuclei are being hyperpolarised by SABRE will also be described.

Polarisation transfer to ^1H

The polarisation of ^1H nuclei by SABRE has been exemplified for a number of applications. These include the hyperpolarisation of biologically relevant materials and the detection of analytes at low concentration or magnetic field, utilising 1D or 2D NMR spectroscopy.

Detection of analytes in complex mixtures or at low concentration

Tessari and co-workers have investigated the use of SABRE to hyperpolarise a complex mixture and, using standard-addition curves, quantify the original amount of analyte present [61]. The complex mixture consisted of sixteen readily polarisable analyte molecules. To exemplify their developed methodology, the original concentrations of nicotinamide, pyrazine, isoxazole and quinazoline were determined using ^1H NMR spectroscopy. These were determined experimentally to have reasonable agreement with the expected amount of analyte present in the original sixteen-component mixture. It must be noted that the original concentration was only an estimate based on an integral comparison between the analyte of interest and the methylene signal of ethanol, which was added as an internal reference. Thus, these data demonstrate that SABRE can be used to not only detect analytes in the micromolar regime but also to quantify them.

A further work by Tessari and co-workers has explored the use of SABRE towards nanomolar detection [59]. In order to achieve this aim, a co-substrate needs to be used based on experimental data which shows that below 25 μM of pyridine, SABRE effectively ceases as all the pyridine ligands are free in solution (a 12.5 : 1 ratio of pyridine : $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ was used). Essentially, solvent molecules (in this instance methanol) compete with the analyte for binding sites on the catalyst, and this effect is more apparent at low concentrations of analyte. 1-methyl-1,2,3-triazole, mtz, was selected as a co-substrate as, in addition to other factors, it binds more strongly to the catalyst than

the solvent and does not possess signals that overlap with that of the analyte pyridine. The use of a co-substrate led to the observation of two new complexes in solution $[\text{Ir}(\text{IMes})(\text{mtz})_2(\text{py})]^+$ and $[\text{Ir}(\text{IMes})(\text{mtz})(\text{py})_2]^+$. A sample consisting of 2 μM of pyridine, 2 mM of $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ and 13 mM of mtz gave a signal-to-noise ratio of 8 : 1 for pyridine after addition of *parahydrogen*, thus establishing the validity of using a co-substrate. A linear dependence was found to exist between SABRE linked signal intensity and substrate concentration in the low micromolar regime. It was, therefore, postulated that detection of 100 nM of pyridine should be possible.

Low-field detection using SABRE

Using a low-field setup, pyridine has been detected down in the nanolitre regime [62]. Pyridine was detected *via* SABRE using a home-built NMR spectrometer (3.9 mT, 166 kHz ^1H resonance) and $[\text{Ir}(\text{COD})(\text{PCy}_3)(\text{py})][\text{PF}_6]$ as the catalyst. Comparison of hyperpolarised ^1H NMR spectra relative to the corresponding thermal ^1H NMR spectra collected following pre-polarisation at 2 T revealed that, through the use of SABRE, 4.9 nL of pyridine could be detected in a single scan. Only when the volume was increased to 4.9 μL could pyridine be detected in the thermal spectrum following pre-polarisation at 2 T. As only 50% *parahydrogen* was used in these experiments this, therefore, indicates a route to improve the signal further, in addition to the limit of detection.

The same low-field setup has been used to detect the drug molecules harmine (harmala alkaloid, used in the treatment of patients with Parkinson's disease), nicotine and morphine in the micromolar regime [63]. For 1 mg of morphine, the hyperpolarised signal is as large as the thermal equilibrium signal that has been pre-polarised at 2 T. Even after dilution by a factor of two, the resulting sample can still be detected due to the hyperpolarised signal produced. Due to the low-field nature of these measurements, it was not possible to fully discern the chemical nature of the nuclei responsible for generating the signal, although tentative assignments were possible. For example, for nicotine, an in-phase doublet was observed following polarisation transfer with a splitting of 0.95 Hz. This equates to a chemical shift difference of 5.7 ppm. Therefore, this chemical shift difference could relate to the aromatic as well as aliphatic regions of the molecule. Similarly, the hyperpolarised spectrum of harmine possessed an anti-phase doublet with a splitting of 0.64 Hz (3.9 ppm). However, tentative assignments are not possible for morphine as the corresponding ^1H NMR spectrum possessed only a single in-phase signal.

Hyperpolarisation of biologically relevant materials

The polarisation of amino acids and some representative small peptides using SABRE has been explored by Appelt and co-workers [64]. A home-built NMR spectrometer was again used to conduct the measurements at 3.9 mT. Using Crabtree's catalyst as the polarisation transfer template, enhanced signals were obtained. In particular, an anti-phase signal was obtained which was ruled not to be due to hydride as it did not possess the $^{2}\text{J}_{\text{HP}}$ splitting recorded in thermal ^1H NMR spectra. High-field ^1H NMR data indicated that glycine was not chemically changed during the SABRE process and is not permanently bound to the catalytic template. Two peptides, GlyGly and ProHisLeu were also polarised by this technique, but highlighted the need for water soluble catalysts as their solubility in d_4 -methanol was poor.

Building on this work, oligopeptides have also been successfully polarised by SABRE [65]. The oligopeptides were effectively SABRE-labelled with nicotinamide or isonicotinamide. The investigated models are shown in Figure 14. For all the investigated oligopeptides, no signal enhancement was observed for the peptide unit. However, the pyridyl motif did polarise with enhancements of 15 – 25-fold being recorded. These decreased to about 10% of their value from moving from d_4 -methanol to a 50 : 50 d_4 -methanol : D_2O mix.

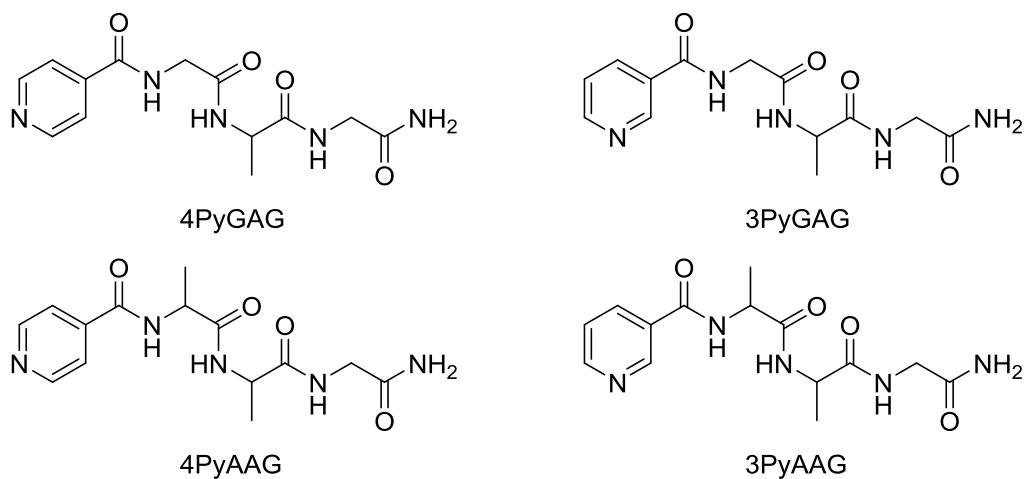


Figure 14. Oligopeptide models investigated by Ratajczyk *et al.* [65].

In a further study, the hyperpolarisation of eNOS (endothelial Nitric Oxide Synthase) substrates were investigated to quantify the production of NO using ^1H NMR spectroscopy [66]. Two generations of eNOS substrates were synthesised; the first generation were focused on the use of a L-arginine carboxylic function to introduce a pyridyl binding group, whereas the second generation made use of N^ω -hydroxy-L-arginine (L-NOHA) to incorporate the pyridyl ring. When these substrates were hyperpolarised, it was generally found that the 4-substituted pyridines produced better enhancements in earth's magnetic field whereas the 3-substituted pyridines were enhanced better at 65 G. The addition of deuterated acetonitrile as a co-ligand, resulted in slightly improved polarisation for the sterically more encumbered substrates. The best enhancement obtained was for the derivative shown in Figure 15, which gave an enhancement of 870-fold for the four pyridyl ring protons after polarisation transfer in earth's magnetic field. This reduced to 322-fold in the presence of d_3 -MeCN. However, for the dipeptide derivatives, no significant enhancement was observed, most likely due to steric hindrance around the binding site. Of the three second generation substrates produced, only the one possessing a guanidyl group showed SABRE derived enhancement and this was small (*ca.* 50-fold following polarisation transfer in earth's magnetic field). It was hypothesised that as the other two second generation substrates possessed a hydroxyl-guanidyl group that, despite being less basic than a guanidyl group, they may bind to the metal in a stronger manner, therefore, leading to partial or full inactivation of the catalyst. All of the substrates led to comparable or higher NO production by eNOS than L-arginine. These substrates could, therefore, be potentially used to monitor NO production in *in vivo*. However, both the signal enhancement and the longevity of the hyperpolarised signal (T_{1s} of the pyridyl protons of the substrate in Figure 15 were only 3.2 – 3.3 s at physiological pH) would need to be improved.

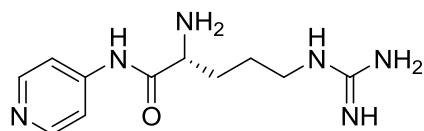


Figure 15. Chemical structure of the pyridyl labelled L-arginine substrate synthesised and polarised by Zanda and co-workers [66].

The tuberculosis drugs pyrazinamide and isoniazid (Figure 16) have been polarised by SABRE in a range of solvents (methanol, d_4 -methanol, ethanol and DMSO) [67]. A Polarizer as described in the section entitled “Automated delivery of SABRE hyperpolarised samples using a Polarizer” was used for these experiments. In addition to the usual setup, a home-built water bath was used to control

the temperature of the system. The enhancement of the three ^1H nuclei in pyrazinamide showed a v-curve dependency with respect to the magnetic-field used for the polarisation transfer. All of these curves indicated that optimum polarisation transfer occurred at 65 G. Of the four solvents used, d_4 -methanol led to the best enhancement being obtained (>60-fold at room temperature). Use of DMSO only led to an enhancement of *ca.* 7-fold and hence is less suitable as a solvent in which to polarise pyrazinamide. Furthermore, polarisation transfer studies using DMSO were only conducted at high temperatures (28.7°C or above), with the maximum enhancement being achieved at 37.5°C. A temperature of 37.5°C also resulted in optimal polarisation transfer for when ethanol was the solvent whereas methanol and d_4 -methanol required a much higher temperature (46.1°C). The latter, however, did yield an enhancement of 1400-fold for the isolated *ortho*-proton after polarisation transfer at 65 G. In the case of isoniazid, polarisation transfer was again optimised at a field of 65 G, where the *ortho*-proton nuclei were enhanced by *ca.* 60-fold. Similar solvent dependence was observed for methanol, ethanol and DMSO as was seen for pyrazinamide. However, for all three solvents the best enhancement was observed at 46.1°C. Use of this temperature to conduct polarisation transfer at 65 G revealed that the ^1H NMR signal of the *ortho*-protons were enhanced by 230-fold. Thus, given these data, pyrazinamide reflects the better spin-system to polarise by SABRE.

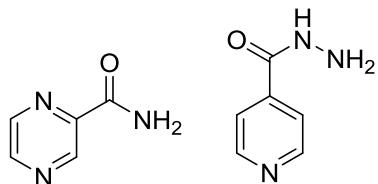


Figure 16. Chemical structures of pyrazinamide (left) and isoniazid (right)

1D and 2D NMR investigations

Quinoline has been the subject of a 1D and 2D NMR investigation, that employs the Polarizer to automate the collection of resulting spectra [68]. The ^1H NMR spectrum of quinoline following polarisation transfer in a 50 G using $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ revealed that polarisation is transferred to the fused secondary ring, as well as the primary ring containing the ligating nitrogen atom. Perhaps more interesting is the fact that some of the ^1H NMR signals on the secondary ring are more enhanced than on the primary ring. The *ortho*-signal on the primary ring was enhanced by 60-fold (signal-to-noise 730 : 1) when detecting 2 μmol of quinoline. Changing the length of time that hydrogen was bubbled through the solution (from 6 s to 20 s) increased the signal enhancement by a further 2.6. Reducing the amount of quinoline further to 0.2 μmol resulted in a single-to-noise ratio of 15 : 1. 20 nmol of nicotinamide could similarly be detected to give a signal-to-noise ratio of 60 : 1. A 1D NOE NMR spectrum of the 2 μmol quinoline sample was collected using four polarisation transfer steps (the steps as outlined in Figure 10 was completed four times). Using a 1.5 s NOE evolution time and exciting the H(4) resonance, led to a cross-peak at the resonant frequency of H(5) with an intensity of 0.9%. Using thermal polarisation, a similar signal-to-noise was achieved after 15000 scans. Furthermore, the thermal spectrum took 20 hours to acquire whereas the polarised spectrum (which includes transfer of the hyperpolarised solution to and from the magnet to the Polarizer) took 81 s. A hyperpolarised 2D OPSY-COSY NMR spectrum was also collected. The spectrum was collected using a *parahydrogen* bubbling time of 2 s and 64 increments, to give a total experiment time of 16 minutes. The incorporation of OPSY resulted in cross-peak encoding being optimised at the start of the COSY sequence thus maximising the observation of cross-peaks over the diagonal. Although the time of this experiment is relatively long, it was also demonstrated that a

hyperpolarised ultrafast COSY could be collected on a 6 μ mol sample of quinoline in a single transient.

A further study also utilised ultrafast methods to obtain 2D spectra in one scan [69]. Initially bipyridine (bpy) was studied using pyridine as a co-substrate. The large enhancement of the *ortho*-pyridyl 1 H NMR signal obscured the β - 1 H of bpy, preventing its facile detection. This was circumvented by using *d*₅-pyridine; this led to the α -protons of bpy increasing in signal strength from 86-fold to 140-fold. An ultrafast 1 H- 1 H COSY spectrum of this solution was acquired and, due to the use of a co-substrate, enabled bpy to be detected at sub-micromolar concentrations. A mixture of five different substrates was then examined. The solution consisted of bipyridine (bpy), pyridine, nicotinamide, *N,N*-dimethylnicotinamide, 2,- 3,- and 4-ethylpyridine in concentrations of the order of 0.5 mM. The ultrafast single scan 1 H- 1 H COSY spectrum obtained, following polarisation transfer is of good quality and enables all but one of the *N*-heterocycles to be detected despite the convoluted nature of the corresponding 1D 1 H NMR spectrum. The *N*-heterocycle not to be detected was 2-ethylpyridine and this was proposed to be due to steric hindrance preventing binding to the catalyst, thus preventing polarisation transfer.

Pyridazine and phthalazine are a further two *N*-heterocycles that have been polarised by SABRE [70]. Both of these substrates display an interesting exchange dynamic; as well as dissociating from the metal catalyst, they also switch ligating atoms (haptotropic shift). Using EXSY, both processes were measured for [Ir(COD(IMes)(pyridiazine)]Cl and [Ir(COD(IMes)(phthalazine)]Cl. The rate of the haptotropic shift is quicker for phthalazine (13.4 s^{-1} compared to 5.8 s^{-1} for pyridazine at 255 K). Following Eyring analysis, the enthalpy of activation for the shift was greater for [Ir(COD(IMes)(pyridiazine)]Cl by 5 kJ mol⁻¹ compared to [Ir(COD(IMes)(phthalazine)]Cl. This represents a stronger Ir-N bond existing in the pyridazine complex. When fully activated in the presence of hydrogen, the two ligands have comparable dissociation rates from the active form of the catalyst. However, they do not possess similar enhancements; a sample consisting of [Ir(IMes)(COD)Cl] (5.2 mM) and 5 equivalents of pyridazine gave a total enhancement of 410-fold whereas the corresponding phthalazine produced only 195-fold. Furthermore, unlike quinoline [68], the remote ring is only weakly polarised (*ca.* 14% relative to the primary ring). The stronger, and therefore shorter, Ir-N bond in the pyridazine system, may result in improved *J*-coupling strength to propagate polarisation relative to the phthalazine system. 1 H MR images of pyridazine and phthalazine were also acquired. The signal-to-noise intensities in these images reflect the signal enhancements observed in the polarised 1 H NMR spectra.

Duckett and co-workers have reported on the synthesis of the COE (cyclooctene) derivatives of [Ir(IMes)(COD)Cl] [71]. Although it was reported that [Ir(IMes)(COE)₂Cl] was unstable (despite forming [Ir(IMes)(H)₂(py)₃][Cl] upon reaction with pyridine and H₂), it was possible to treat [IrCl(COE)₂]₂ with pyridine and H₂ in *d*₈-THF which gave IrCl(H)₂(COE)(py)₂. The chemical structure of this complex is shown in Figure 17. The pyridine dissociation rate was found to $7.8 \pm 0.1\text{ s}^{-1}$ at 298 K; at this temperature the hydride exchange rate into H₂ could not be determined. However, the two distinct hydride ligands interchanged at a rate of $3.6 \pm 0.1\text{ s}^{-1}$ whilst the COE CH proton sites interchanged at a rate of $3.8 \pm 0.1\text{ s}^{-1}$. Changing the concentration of pyridine and H₂ did not affect these ligand rates. In the SABRE enhanced 1 H NMR spectrum, the hydride, COE and pyridine ligands are all enhanced. The enhancement of pyridine was shown to be dependent on the concentration of pyridine; an 8-fold excess led to a 210-fold signal gain of the *ortho*-hydrogen of pyridine, whereas using a 5.6-fold excess, the enhancement increased to >500-fold. After a period of 24 hr, a new complex is observed to form due to the appearance of three new hydride signals. This complex

proved to be $(Cl)(H)(Py)Ir(\mu-Cl)(\mu-H)(\kappa-\mu-NC_5H_4)Ir(H)(py)_2$ and as such contained four distinct pyridine ligands. This complex is SABRE inactive but does catalyse H-D exchange.

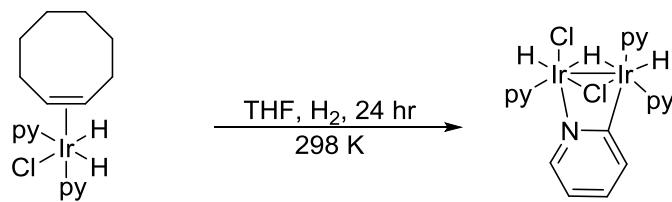


Figure 17. Reaction scheme for the formation of the SABRE-inactive binuclear complex from the mononuclear SABRE active complex prepared by Duckett and co-workers [71].

Polarisation transfer to ^{13}C

Polarisation transfer to ^{13}C via SABRE was first exemplified using the $[Ir(PCy_3)(COD)(py)]$ template [23]. 6 nmol of pyridine were polarised in a 2×10^{-2} T magnetic field to afford the $^{13}C\{^1H\}$ NMR spectrum in a single scan. The *meta*-signal is the most intense, with the *ortho*- and *para*-signals being of similar intensity. In a separate study 50 μ mol of nicotinamide was polarised and a refocused $^{13}C\{^1H\}$ NMR spectrum collected. All six resonances were observed in a single transient, with the quaternary carbon of the amide being the least intense.

In order to determine a strategy to determine polarisation transfer to the two ^{13}C sites of acetonitrile, different isotopomers were polarised via SABRE and the results reported [72]. Deuterated acetonitrile resulted in no visible enhancement of either the ^{13}C site due to the absence of 1H - 1H coupling between the *parahydrogen* derived hydrides and that of the substrate during polarisation transfer. Labelling the methyl group of acetonitrile with a ^{13}C label only resulted in a 14-fold enhancement gain. The large $^1J_{HC}$ (~ 145 Hz) is not commensurate with the J_{HH} coupling that exists between the *parahydrogen* derived hydride and the proton nuclei of acetonitrile. Thus, it is hard to propagate polarisation to this site. However, ^{13}C -labelling the nitrile group results in significant signal gains. Following polarisation transfer at earth's field, a single transient ^{13}C spectrum yields a signal-to-noise value of 105 for the quaternary carbon, which possesses anti-phase character. Furthermore, an antiphase signal is observed for acetonitrile ligands bound to the catalyst *trans* to hydride. In addition, it was shown that the ^{13}C enhancement and appearance over a range of polarisation transfer fields (earth's to 140 G) could be modelled theoretically and was in excellent agreement with experimental results.

The ^{13}C nuclei of nicotinamide have been the subject of a study in order to investigate ^{13}C polarisation [48]. Using the Polarizer (as detailed in the section entitled “Automated delivery of SABRE hyperpolarised samples using a Polarizer”), nicotinamide was polarised by setting the magnetic field to 70 G and with the magnetic field switched off. Four-scan averages obtained revealed that polarisation transfer was superior in earth's magnetic field (magnetic field switched off). The signals were suitably more enhanced compared to polarisation transfer at 70 G. However, these spectra could be improved by using ^{13}C INEPTnd; in this instance polarisation transfer occurred at 70 G to optimally polarise the 1H nuclei in order to maximise the signal intensity of the ^{13}C sites following INEPT transfer. An evolution period of 12 Hz was used during the pulse sequence to optimise transfer further. All of the signals manifested in the ^{13}C spectrum presented as anti-phase signals.

The hyperpolarised single scan ^{13}C spectra of quinoline enables all nine of the expected signals to be observed, although the carbon located at the C(2) position (δ 150.8) is significantly reduced [68]. In addition, the detected resonances appear with a degree of antiphase character. To improve the

quality of the spectrum, and also the signal intensity of C(2), INEPT was employed. A refocused single scan $^{13}\text{C}\{\text{H}\}$ INEPT NMR spectrum of 60 μmol of quinoline acquired following polarisation transfer at 50 G resulted in a signal-to-noise ratio of 40 : 1 for C(2). To achieve the same signal-to-noise ratio using thermal polarisation would take 4096 scans. Using the Polarizer, a $^{13}\text{C}-^1\text{H}$ HMBC spectrum was collected on 2 μmol of quinoline. The 32 increment spectrum was acquired using a single scan per increment, and from which, all nine of the ^{13}C signals were located. When the sample was polarised in earth's magnetic field the cross peaks derived from short range couplings, whereas when polarised at 60 G, the cross peaks derive from long-range couplings.

Polarisation of the aromatic ^{13}C sites of ^{15}N -labelled pyridine has also been probed *via* conventional SABRE and SABRE-SHEATH (readers are referred to the SABRE-SHEATH section for full details of the method) [73]. Conventional SABRE, following polarisation transfer at 6 mT, yields three antiphase signals. Using SABRE-SHEATH, these signals manifest as in-phase signals. However, the enhancements obtained were marginal; they were 7-fold or less. The signals obtained using conventional SABRE were visibly more enhanced. However, a later report optimised the conditions for transfer such that an enhancement of *ca.* 1900-fold was obtained at 9.4 T, which equates to 1.5% polarisation [74]. The magnitude of the enhancement was highest for the *ortho*-carbon site and lowest for the *para*-carbon site, which mirrored the size of the J_{HC} couplings between each site and that of the hydride. The relatively short ^{13}C T_1 s at both earth's field and high-magnetic field (*ca.* 5-8 s) reduces the enhancement observed considerably. The conditions optimised included the temperature (maximum at 25°C), the magnetic field in which polarisation was conducted (two maxima at 0.5 μT and -0.75 μT) and the flow rate at which the 7 atm. of *parahydrogen* was delivered. The efficiency of polarisation transfer was also probed. It was found that the presence of ^{14}N instead of ^{15}N in the pyridine significantly reduced the efficiency of SABRE-SHEATH. As the *ortho*-carbon signal was reduced, it was concluded that the quadrupolar ^{14}N relaxation acts to as a sink for hyperpolarisation.

Polarisation transfer to ^{15}N

The effect on polarisation transfer using ^{15}N -labelled substrates was first exemplified using $[\text{Ir}(\text{COD})(\text{PPh}_3)_2]\text{BF}_4$ [25]. In the presence of pyridine, $[\text{Ir}(\text{H})_2(\text{PPh}_3)_2(\text{py})_2]^+$ and $[\text{Ir}(\text{H})_2(\text{PPh}_3)(\text{py})_3]^+$ were formed in methanol and were found to possess hydride resonances at δ -21.64 and δ -21.80 respectively. When ^{15}N -labelled pyridine was employed, the hydride resonances of both complexes were enhanced. Direct polarisation transfer to ^{15}N was only reported to occur when specialised heteronuclear polarisation transfer experiments were utilised. Without them, no ^{15}N signals of bound ^{15}N -labelled pyridine were visible after 1024 scans. The most efficient transfer of magnetisation was achieved using the PH-INEPT pulse sequence. This resulted in a signal-to-noise ratio of 27.9. The PH-INEPT pulse sequence was found to be superior to the PH-INEPT⁺ experiment, as relaxation during the refocusing period led to only a third of the potential signal intensity surviving to the point of observation. The PH-INEPT⁺ pulse sequence yielded a signal-to-noise of 10.5.

A further study utilised a zero-field magnetometer in which to detect the polarised signals of ^{15}N -pyridine following polarisation transfer by Crabtree's catalyst [75]. A DC magnetic-field pulse was used to convert the spin order involving heteronuclear pairs in to observable magnetisation. A 40 mM solution of hyperpolarised ^{15}N -pyridine possessed a signal-to-noise ratio of 300. This contrasted with neat ^{15}N -pyridine (14 M), which after the collection of 128 transients only produced a signal-to-noise ratio of 70. A sensitivity gain of $\sim 1.7 \times 10^4$ was, therefore, obtained using SABRE.

LIGHT-SABRE

The initial experiments to relating to the polarisation transfer to ^{15}N have been significantly improved by employing LIGHT-SABRE (Low-Irradiation Generation of High Tesla SABRE) [76]. The coherent transfer of polarisation to the metal-bound ^{15}N -labelled pyridine was shown for the complex $[\text{Ir}(\text{IMes})(\text{py})_3(\text{H}_2)]^+$. A continuous wave (CW) was applied to the bound ^{15}N -labelled pyridine resonance; the power of the pulse matched the sum (or difference) of the hydride-hydride and ^{15}N - ^{15}N *J*-coupling (see Figure 18 for sequence). This lead to the transfer of polarisation of magnetisation from the *parahydrogen* derived hydrides to ^{15}N that manifested itself as x-magnetisation. However, due to exchange processes, this x-magnetisation would diphase immediately following dissociation from the metal centre of the catalyst. Thus, selective 90° pulses were incorporated into the LIGHT-SABRE pulse sequence and applied n times, at a repetition rate comparable to the exchange rate and for a time comparable to T_1 . Optimization revealed that values of $n = 15$ and $\tau_{\text{CW}} = 0.5$ s should be used (total pulse length of 7.5 s). When these conditions were employed, a sample consisting of 63 mM of ^{15}N -labelled pyridine in methanol was enhanced by 150-fold relative to a 128-scan thermal spectrum. Reducing the concentration of ^{15}N -pyridine to 2.5 mM increased the enhancement further to 480-fold. In both instances, the ^{15}N -pyridine to catalyst concentration was 10:1.

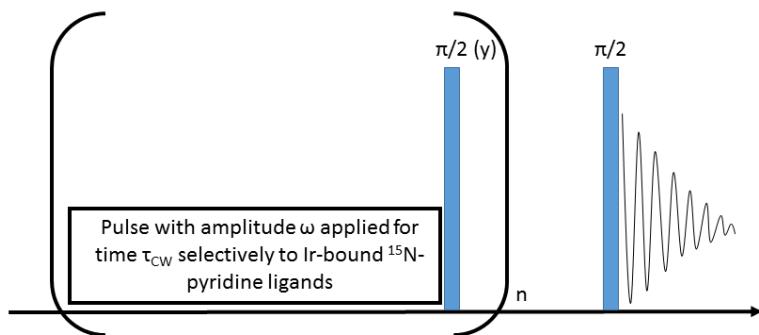


Figure 18. LIGHT-SABRE pulse sequence. The amplitude of the pulse, ω , is $2\pi(J_{\text{HH}} \pm J_{\text{NN}})$ and τ_{CW} is $1/V2\Delta\Delta_{\text{NH}}$

SABRE-SHEATH

LIGHT-SABRE enables SABRE to be conducted in the magnet and under constant bubbling of *parahydrogen*. This means that the method is highly reproducible, as it eliminates experimental uncertainties resulting from sample transfer and discontinued *parahydrogen* supply. Nonetheless, a field cycling method was found to be far superior in yielding significant polarisation values of ^{15}N . SABRE-SHEATH (SABRE in SHield Enables Alignment Transfer to Heteronuclei) [77] utilises a μ -metal shield to conduct the hyperpolarisation within for approximately 30 s. Following hyperpolarisation transfer, the sample is transferred rapidly (~ 4 s) to the NMR spectrometer for detection. At concentrations of 4 mM of ^{15}N -labelled pyridine and 0.24 mM $\text{Ir}(\text{IMes})(\text{COD})\text{Cl}$, a 30,000-fold enhancement for free ^{15}N -labelled pyridine was achieved. This equates to a polarisation value of 10%. Another study that used 70 mM of ^{15}N -labelled pyridine and 2 mM $\text{Ir}(\text{IMes})(\text{COD})\text{Cl}$ yielded a 2500-fold enhancement [60]. Crucially, the signal manifests as a purely emissive in-phase triplet. This contrasts to the anti-phase signal observed in the corresponding ^{15}N NMR spectrum when traditional SABRE is conducted in a polarisation transfer field of ~ 6 mT along with a lower enhancement (only 460-fold). The appearance of the signal is important from an imaging application perspective because broadened lines can lead to partial signal cancellation. Furthermore, employment of SABRE-

SHEATH enables not just the detection of free ^{15}N -labelled pyridine, but also equatorially bound (*trans* to hydride) and axial (*trans* to IMes) ^{15}N -labelled pyridine. This is noteworthy because the latter is not observed under traditional SABRE conditions due to the anti-phase nature of the signal leading to signal cancellation. ^{15}N -labelled nicotinamide was also polarised employing SABRE-SHEATH; a 48 mM solution yielded an enhancement of 20,000-fold over the thermal reference, which equates to a polarisation of 7%. Compared to ^{15}N -labelled pyridine for which a higher concentration (63 mM) yielded an enhancement of only 3000-fold, this represents a more biologically relevant concentration and as such could be used in MRI studies.

A further report on SABRE-SHEATH delved into the physical constraints that affect polarisation transfer. *Parahydrogen* pressure, flow rate, sample temperature, catalyst-to-substrate ratio, T_1 relaxation, and reversible oxygen quenching were all studied using a test system of ^{15}N -labelled pyridine and $\text{Ir}(\text{IMes})(\text{COD})\text{Cl}$ in d_4 -methanol [73]. The exposure of a sample to air was shown to have a very dramatic effect on the enhancement observed, however, this situation was recoverable upon exposure to *parahydrogen* for long time periods. A SABRE-SHEATH ^{15}N NMR spectrum produced after 20 min of *parahydrogen* bubbling, (using a 20% duty cycle i.e. bubbling turned on for 30 s and then off for 2 min) yielded an enhancement of 3600-fold for free ^{15}N -pyridine (48 mM). Exposure to air for \sim 5 s resulted in the enhancement dropping to 185-fold. However, the 3600-fold enhancement was fully recovered within 8 min following continuation of *parahydrogen* bubbling using the same duty cycle as before. The T_1 of free ^{15}N -pyridine was found to differ considerably; in the magnetic shield ^{15}N T_1 was 10.4 ± 0.9 s, at \sim 6 mT 28.4 ± 2.6 s and at 9.4 T 68.6 ± 2.4 s. The lowest T_1 is observed in the magnetic shield, the magnetic field in which polarisation transfer occurs in SABRE-SHEATH. Thus, T_1 modulates the build-up rate and the maximum polarisation attainable. Understanding these conditions enabled the collection of a ^{15}N image of ^{15}N -labelled pyridine (20 mM), the first such image produced by *parahydrogen*-based hyperpolarisation. The image produced, which took 0.4 s to acquire, had a $2 \times 2 \text{ mm}^2$ spatial resolution and the hyperpolarised liquid was visible within a 6.6 mm internal diameter high-pressure HPLC column.

The hyperpolarisation of neat liquids has been demonstrated using SABRE-SHEATH [78]. Using [$\text{Ir}(\text{IMes})(\text{COD})\text{Cl}$] as the pre-catalyst, the ^{15}N of neat pyridine was enhanced by up to 2900-fold. This equates to roughly 1% polarisation. Interestingly, the *ortho* ^1H signal of pyridine was only enhanced 4-fold using conventional SABRE. For these experiments, *parahydrogen* was bubbled into the tube containing the catalyst and neat pyridine. The ^{15}N signal exhibited a strong dependence on the flow rate of *parahydrogen* over the range studies (0 – 150 scc/min). The signal enhancement obtained was independent of the pressure of *parahydrogen* supplied during polarisation transfer. The authors note that the low concentration of *parahydrogen* in solution relative to the catalyst, and also the significantly shorter T_1 in the micro-Tesla field in which polarisation transfer is conducted relative to the 9.4 T T_1 (5.5 ± 0.5 and 60.8 ± 0.6 s respectively), results in a smaller ^{15}N signal enhancement than what could be achieved. This was exemplified by using neat ^{15}N labelled pyridine; the ^{15}N signal was enhanced only 33-fold, which is two orders of magnitude different to the natural abundance sample. This is despite an increase in T_1 within the micro-Tesla field (10.2 ± 1.1 s).

SABRE-SHEATH has been utilised to hyperpolarise a wide range of ^{15}N containing molecules [79]. Importantly, molecules that did not contain a pyridyl motif were exemplified using this approach. A selection of the molecules that were polarised and their enhancement levels at 8.5 T are shown in Figure 19. These molecules suggest that sp - and sp^2 -hybridised ^{15}N sites can be hyperpolarised. In addition, it is prudent to note that the ^{15}N site of benzonitrile was polarised (21000-fold enhancement) whereas it had been previously demonstrated that the ^1H nuclei could not [72]. T_1 values were obtained for all the substrates; values in the order of minutes were obtained for

acetonitrile, benzonitrile, CHCA (α -cyano-4-hydroxycinnamic acid), nicotinamide and a representative Schiff base. Other Schiff bases have been polarised using either the $[\text{Ir}(\text{IMes})(\text{COD})(\text{py})]\text{PF}_6$ or $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ templates; smaller Schiff bases produced bigger enhancements when polarised by the latter, whereas *vice versa* is true for larger Schiff bases [80]. The ^{15}N T_1 times obtained in ref. [79] were strongly dependent on the magnetic field and molecular structure. The reason for this was ascribed to chemical shift anisotropy (CSA) which scales with B_0^2 . As a result of this study, acetonitrile was chosen as a substrate to be optimised. The maximum enhancement was obtained using 0.25 mM $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ and 5 mM $^{15}\text{N}-\text{CH}_3\text{CN}$ in the presence of 0.25 mM pyridine as a stabilizing co-ligand. This resulted in $P_{15\text{N}} = 7\%$. The polarisation value was further optimised by increasing the flow rate of hydrogen supplied to the reaction vessel; a flow-rate of 130 sccm/min rather than 80 sccm/min increased the polarisation to 16%.

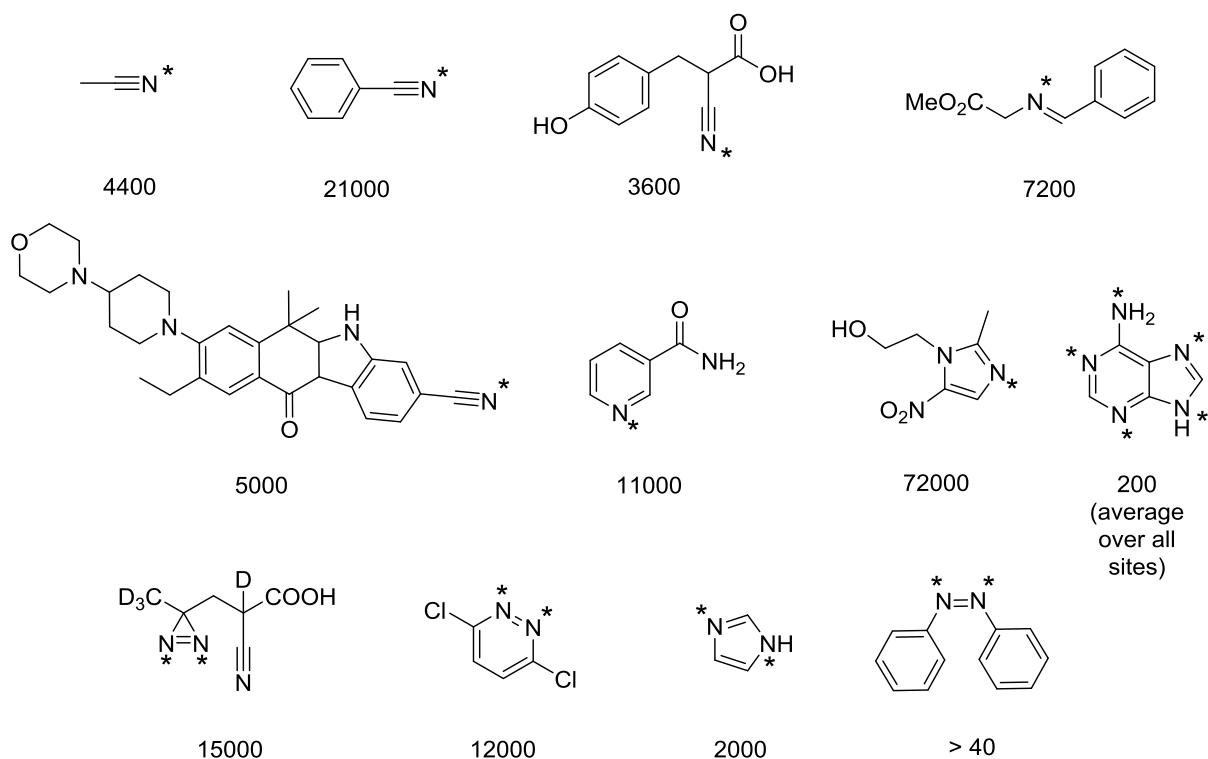


Figure 19. ^{15}N containing molecules polarised by Theis *et al.* ^{15}N labels are with a *. The ^{15}N enhancements obtained at 8.5 T relative to thermal measurements are indicated underneath each molecule.

An improved synthesis for ^{15}N -labelled nicotinamide, has been reported that makes use of the Zincke salt [81]. The nicotinamide is 98% isotopically pure with respect to the ^{15}N label. When polarised using $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ and employing SABRE-SHEATH, an enhancement of 7,300-fold was obtained in d_4 -methanol. This value is comparable to those obtained using the 66% isotopically pure ^{15}N -nicotinamide [77, 79]. However, using the 98% isotopically pure ^{15}N -nicotinamide results in almost 50% more hyperpolarised ^{15}N magnetisation, which is an advantage for imaging purposes.

Polarisation of metronidazole (^{15}N at natural abundance) *via* SABRE-SHEATH has yielded substantial enhancements [82]. A 72,000-fold enhancement was observed in the ^{15}N NMR spectrum ($P = 24\%$) for 50 mM of metronidazole in d_4 -methanol. Only 80% *para*hydrogen was used in the experiment, so a polarisation of 32% is feasible if 100% *para*hydrogen was used. It should be noted that metronidazole possesses *ortho*-substituents and these do not prevent successful polarisation of the molecule, unlike the picolines and lutidines noted elsewhere [78].

Imidazole has been polarised using SABRE-SHEATH with a view towards pH sensing due to the sizable chemical shift difference (~30 ppm) that occurs upon binding to the SABRE catalyst [83]. The imidazole (0.1 M) ^{15}N nuclei were polarised by *ca.* 2000-fold when measured at 9.4 T. Only 50% *parahydrogen* was used in this measurement and so an enhancement of 6000-fold is theoretically possible. Of significant note, however, is the fact that a 2000-fold enhancement was obtained in a 50:50 d_4 -methanol:aqueous (pH 12) mix. The pH of the solution was varied so that it was above and below the pK_a (pK_a of imidazole ~7.0) and this showed that the chemical shift of the imidazole nitrogen nuclei moved from 175 ppm at pH 1 to 205 ppm at pH 12. The biggest change occurred in the 6.5 to 7.5 pH region; here a resolution of 0.1 unit of pH per 1.5 ppm of ^{15}N shift was observed. The presence of imidazole rings in peptides and proteins coupled with the ^{15}N chemical shift change and associated increases in signal-to-noise afforded by SABRE-SHEATH, means that this system has potential for pH sensing over the physiological pH range.

Further research on SABRE-SHEATH has focused on utilising the technique in aqueous media [84]. The water soluble pre-catalyst, $[\text{Ir}(\text{IDE})\text{(COD)}\text{Cl}]$ [85], was employed to ensure compatibility with the D_2O and H_2O solvents used for the study. A range of substrates were again selected to probe the polarisation transfer to ^{15}N . ^{15}N -Pyridine, ^{15}N -benzonitrile, ^{15}N -CHCA, 2-cyano-3-(d_3 -methyl- $^{15}\text{N}_2$ -diazirine)-propionic acid and ^{15}N -nicotinamide were all successfully polarised. Interestingly, the normalised hyperpolarised signals as a function evolution magnetic field strength revealed that the curve for CHCA polarised using $[\text{Ir}(\text{IDE})\text{(COD)}\text{Cl}]$ showed a similar response to those obtained for acetonitrile polarised by $[\text{Ir}(\text{IMes})\text{(COD)}\text{Cl}]$.

2-cyano-3-(d_3 -methyl- $^{15}\text{N}_2$ -diazirine)-propionic acid has been the subject of a further SABRE-SHEATH study [86]. DFT calculations revealed that the diazirine ligated through a η^1 arrangement through one of the nitrogen atoms, rather than a $\eta^2\text{-N=N}$ attachment. The use of $[\text{Ir}(\text{COD})(\text{IMes})(\text{py})]\text{[PF}_6]$ instead of $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ resulted in a factor of ~3-fold improvement in the hyperpolarisation levels. The complex $[\text{Ir}(\text{IMes})(\text{diazirine})_2(\text{py})]\text{[PF}_6]$ is formed after activation, in which the pyridine is *trans* to IMes and therefore avoids the pyridine competing for sites *trans* to hydride and thus acting as a polarisation acceptor. Following hyperpolarisation, both z-magnetisation and singlet order were produced. A 15,000-fold enhancement (5% polarisation) was observed for the z-magnetisation produced with a T_1 of 5.8 ± 0.3 minutes at 12 mT, whereas the singlet order gave a 5000-fold enhancement (3% polarisation) and a T_s of 23 ± 3 minutes at 0.3 mT. The former was produced under SABRE-SHEATH conditions, whereas the latter, which relies upon $J_{\text{HH}} = \pm J_{\text{NN}}$ is not dependent on magnetic field provided that the singlet state remains close to a eigenstate of the Hamiltonian. Well-suited lifetimes for the singlet state were found at fields between 0.5 and 200 G. It should be noted that the singlet readout was only ~60% efficient, and so the polarisation of z-magnetisation and singlet order are comparable, although the build-up rate for singlet order is longer than z-magnetisation (T_b for the former is 5 min compared with 2 min for the latter). The section entitled “Generation and utilisation of long-lived states by SABRE” details further the generation and use of singlet states in conjunction with SABRE.

Polarisation transfer to ^{19}F

To date, only one analyte that contains fluorine has been polarised by SABRE; 3-fluoropyridine. This analyte is largely analogous with pyridine, and as J_{HF} couplings tend to be of the same order of magnitude as J_{HH} couplings, it is therefore not surprising that this molecule can be polarised successfully. The hyperpolarised ^{19}F signal presents as a complicated anti-phase multiplet that is enhanced relative to the trace at thermal equilibrium [23]. $[\text{Ir}(\text{COD})(\text{PCy}_3)(\text{py})]$ was used as the polarisation transfer catalyst. Pentafluoropyridine has been the subject of a SABRE study but did not yield any polarisation using either $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ or Crabtree’s catalyst as the polarisation transfer

templates [87]. Investigating ^{19}F containing substrates as analytes for SABRE should be a future focus within this research field, especially given the number of drugs that possess ^{19}F nuclei (in 2010 it was calculated that about 20% of administered drugs contained fluorine atoms or fluoroalkyl groups [88]) and that in MRI there is no competing background signal.

Polarisation transfer to ^{31}P

Phosphorous is a component of many ligands that are used in inorganic catalysis.

Triphenylphosphine, PPh_3 , is one such example. Zhivonitko *et al.* have shown that PPh_3 can be polarised using $[(\text{PPh}_3)_3\text{Ir}(\text{H}_2)\text{Cl}]$ [36]. The complex possesses two unique phosphorous environments, as shown in Figure 20. Interestingly, after polarisation transfer to a heated sample (60°C) inside a μ -metal shield (*ca.* 1 μT), polarisation transfer was observed to both bound PPh_3 ligands and also free PPh_3 . ^{31}P NMR signal enhancements of *ca.* 120-fold were obtained in a 7 T measurement field. Heating the sample to 80°C gave higher enhancements of *ca.* 260-fold. The reason for the increase is due to an increase in the exchange rates of both hydride and phosphine by 3 – 7-fold. EXSY data showed that the phosphines *trans* to one another in $[(\text{PPh}_3)_3\text{Ir}(\text{H}_2)\text{Cl}]$ do not exchange within the relaxation time of ^{31}P magnetisation.

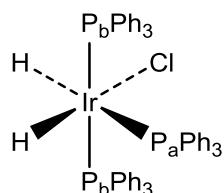


Figure 20. Chemical structure of $[(\text{PPh}_3)_3\text{Ir}(\text{H}_2)\text{Cl}]$ showing the two unique phosphorous environments (labelled a and b)

A mixed NHC and PPh_3 system has also yielded polarisation transfer to ^{31}P via SABRE. The complex shown in Figure 21 yielded a 30-fold ^{31}P signal enhancement [89]. The acetonitrile ligands were also shown to polarise in the ^1H NMR spectrum. In the analogous complex in which one of the acetonitrile ligands is exchanged for pyridine, the methyl proton nuclei of acetonitrile were enhanced by *ca.* 23-fold following polarisation transfer at 20 G. The total pyridine ^1H NMR enhancement was 35-fold, but to obtain this maximum, a polarisation transfer field of 140 G was required. The enhancement of pyridine increased further when both acetonitrile ligands were replaced by pyridine.

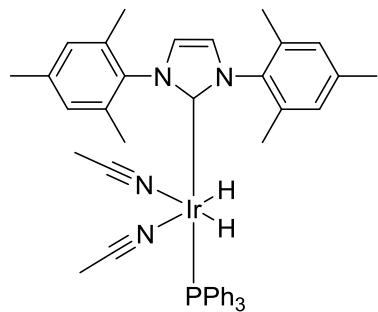


Figure 21 Chemical structure of $[\text{Ir}(\text{IMes})(\text{PPh}_3)(\text{MeCN})_2(\text{H})_2]^+$ utilised by Fekete *et al* [89]

Further work on phosphorous containing compounds, in terms of being polarised by SABRE, led to the synthesis of pyridine derived molecules possessing phosphonate groups. This work demonstrated that the *ortho*-site of the pyridine ring was not a suitable position for tethering groups to as it prevented polarisation transfer. This was exemplified for 2-pyridyldiethylphosphonate; no ^{31}P polarisation was observed and only a marginal enhancement of 2-fold in the ^1H NMR spectrum. Similar results have been demonstrated for 2-picoline and 2,6-lutidine [78]. In contrast, substitution

at the *meta*- and *para*-position were amenable to polarisation transfer. Total ^1H NMR enhancements of 1499-fold were achieved at a polarisation transfer field of 45 G for 4-pyridyldiethylphosphonate. The ^{31}P NMR revealed a 545-fold enhancement. It is noteworthy that the analogous phosphine, phosphine oxide and phosphine sulphide did not yield as substantial enhancements. However, tethering the diethylphosphonate to the *meta*-position resulted in improved polarisation transfer; the ^1H nuclei are enhanced to a total of 2866-fold whereas the ^{31}P signal enhancement was enhanced by 336-fold. The latter equates to a polarisation value of 0.44%. This value was increased by further substituting both *meta*-positons with diethylphosphonate groups. The ^1H and ^{31}P signals were enhanced by 3689- and 860-fold respectively. Interestingly, it was noted that a lower number of equivalents of this ligand could be used in this instance relative to the other substrates studied (6 equivalents instead of 7 equivalents). This is because the increased steric bulk of the di-*meta*-substituted diethylphosphonate results in a chloride occupying one of the coordination sites *trans* to hydride in the active catalyst instead of a substrate molecule. This arrangement results in improved polarisation transfer to the only di-*meta*-substituted diethylphosphonate ligand that is *trans* to hydride. Deuterating the ethyl groups of di-*meta*-substituted diethylphosphonate ligand further improved the ^{31}P hyperpolarised signal. Following polarisation transfer at 0.5 G a ^{31}P NMR signal gain of 3588-fold (2.3% polarisation) was achieved. A diminished signal gain of 2251-fold (1.4% polarisation) was obtained after polarisation transfer at 45 G, but this signal possessed significant in-phase character. This observation was, therefore, prudent to conducting imaging studies which are detailed in Imaging Applications of SABRE section.

Polarisation transfer to ^{29}Si and ^{119}Sn

^{29}Si and ^{119}Sn are two other nuclei that have been shown to be polarisable *via* SABRE [90]. The two molecules synthesised and analysed by SABRE are shown in Figure 22. A thermal spectrum of both molecules were collected; 5-(tributylstannyl)pyrimidine required 3096 averages and 5-(trimethylsilyl)pyrimidine required 128 averages to give spectra with appropriate signal-to-noise. For the former, enhancements of 803 ± 73 -fold (H-2 resonance) and 1486 ± 156 -fold (H-4 and H-6 resonances) were observed in the ^1H NMR spectrum after polarisation transfer at 70 G.

[Ir(IMes)(COD)Cl] was employed as the SABRE pre-catalyst. To optimise transfer to ^{119}Sn , the excess amount of substrate relative to SABRE catalyst was modulated in order to deliver the maximum enhancement. A 17-fold excess of substrate was required to deliver a SNR of 375 : 1 in the fully coupled ^{119}Sn NMR spectrum and 1099 : 1 in the $^{119}\text{Sn}\{^1\text{H}\}$ NMR spectrum (decoupling applied only to butyl protons). This gave enhancements of 687- and 772-fold respectively. The use of [Ir(SIMes)(COD)Cl] resulted in approximately 50% lower enhancements due to the rate of H_2 loss from the active complex being 2.29 s^{-1} compared to 0.56 s^{-1} when [Ir(SIMes)(COD)Cl] is employed. Incorporation of SIMes into the catalyst structure therefore promotes more rapid ligand loss, which is detrimental in terms of polarisation transfer. Complexes formed with 5-(trimethylsilyl)pyrimidine yielded the same trend, and this was reflected in the IMes catalyst producing the largest enhancement. The hyperpolarised non-decoupled ^{29}Si NMR spectrum gave a SNR of 200. Furthermore to the hyperpolarisation study, both substrates form tri-nuclear complexes with the general formula $[\text{Ir}(\text{H})_2\text{Cl}(\text{NHC})(\mu\text{-pyrimidine-}\kappa\text{N: }\kappa\text{N}')]_3$.

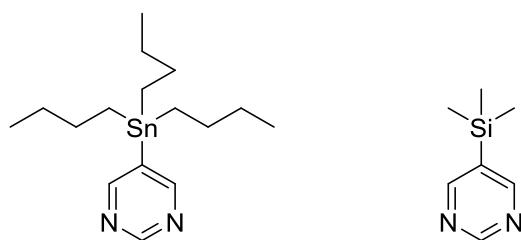


Figure 22 Chemical structures of 5-(tributylstannyl)pyrimidine and 5-(trimethylsilyl)pyrimidine

Towards performing SABRE in water-based solvent systems

It is desirable to conduct SABRE experiments in water as this enables biologically relevant molecules, such as proteins and amino acids, to be studied more readily. The production of water soluble catalysts also aids the collection of *in vivo* images as physiological pH can be more easily maintained. Furthermore, it avoids the use of solvents that are incompatible for *in vivo* use, such as methanol.

A method for achieving 1% NMR polarisation in water in less than one minute using SABRE has been reported for 3-amino-1,2,4-triazine (ATZ) [91]. A sample consisting of ATZ and [Ir(IMes)(COD)Cl] was dissolved in methanol and hydrogen bubbled through the solution for at least 2 hours. Then water was added. Following evaporation under reduced pressure for 10 minutes, the resulting molar fraction of methanol : water was 10 : 90. This solution was used for SABRE-based measurements. Following polarisation transfer at 65 G and at 54.4°C for 30 s, an enhancement of 170 ± 45 -fold was obtained for the two protons of ATZ. The solvent signals were suppressed using excitation sculpting with gradient dephasing. Interestingly, using this methodology, no SABRE enhancement was observed for pyridine, isoniazid and pyrazinamide. This is despite the enhancement of ATZ in methanol being similar to that of the aforementioned analytes.

In a separate study, it was found that following activation of [Ir(IMes)(COD)Cl] in the presence of pyridine or nicotinamide and H₂ in ethanol or methanol, the generated complex could be dried and then dissolved in pure water [51]. Using this approach, the four aromatic proton environments of nicotinamide were enhanced to give a total enhancement of 100-fold following polarisation transfer at 6 ± 4 mT. Although a decrease from the *ca.* 230-fold enhancement observed in *d*₆-ethanol, this method again shows that SABRE can be conducted in aqueous media.

The design of new water soluble catalysts has also been the focus of some research groups. Fekete *et al.* modified their mixed phosphine / NHC system [89] so that the phosphine ligands possessed water solubilising sulfonated groups [92]. Complexes of the type [Ir(IMes)(py)(H)₂(MeCN)(L)]BF₄ were produced in the presence of H₂, MeCN and pyridine, where L = PPh₃ or one of the phosphine ligands in Figure 23. Although these complexes are readily soluble, they failed to catalyse the SABRE of pyridine in aqueous solvent media. In an extension to this study, a series of water-soluble azonium salts were prepared. These were superior to the aforementioned mixed NHC / phosphine complexes; the pre-catalyst as seen in Figure 24 resulted in a 750-fold enhancement of the ¹H nuclei of pyridine in *d*₄-methanol after polarisation transfer at 60 G. However, when moving towards aqueous media (67% D₂O, 3% DMSO, 30% ethanol), only a 9.4-fold enhancement was obtained for pyridine. It is noteworthy that DMSO was included to improve the hydrogen solubility of H₂ in solution and that it was not polarised in the process. Polarisation of 4-hydroxypyridine and oxazole resulted in enhancement gains of 22- and 2.7-fold in 70% D₂O : 30% ethanol respectively.

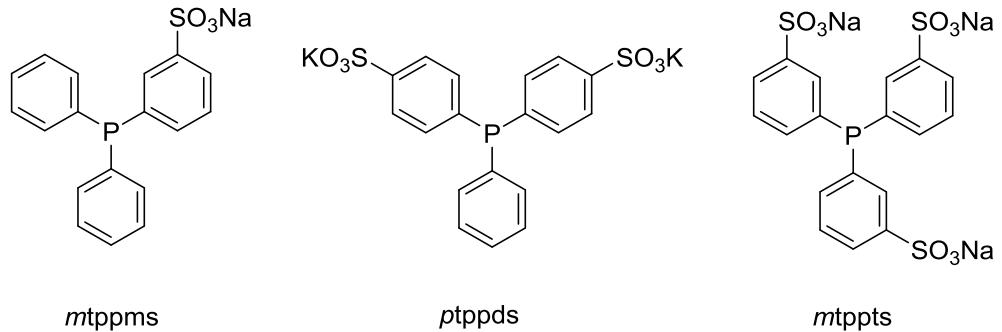


Figure 23. Phosphine ligands used to create water soluble catalysts (*mtpcms* =mono-*meta*-sulfonated triphenylphosphine, *ptpdds* = di-*para*-sulfonated triphenylphosphine and *mtppts* = *meta*-tri-sulfonated triphenylphosphine)

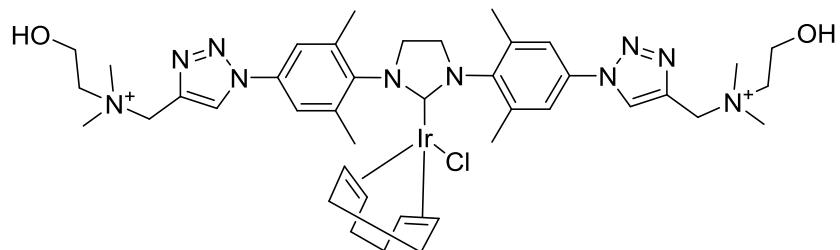


Figure 24. Pre-catalyst synthesised by Fekete *et al* [92].

Spannring *et al.* have modified the NHC of the pre-catalyst $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ so that it possesses water solubilising groups [85]. Exchanging the mesityl groups for 3,4,5-tri(methoxy)benzyl groups to produce the NHC Itome, resulted in a catalyst that was insoluble in D_2O . However, exchanging the mesityl groups for 3,4,5-tri(diethyleneglycol)benzyl groups did result in a catalyst that was soluble in water and D_2O . The structure of this catalyst, $[\text{Ir}(\text{IDEG})(\text{COD})\text{Cl}]$, is shown in Figure 25. Following polarisation transfer at 80 G, total enhancements of 64-fold were observed for pyridine in a measurement field of 300 MHz. A 48 hour activation period was required prior to the measurement. Methyl nicotinate, *N*-methyl nicotinamide and nicotinamide were also polarised yielding similar enhancements. Crucially this system does not require an organic solvent in the activation or subsequent steps. In addition, the low solubility of H_2 does not preclude SABRE from occurring in pure water.

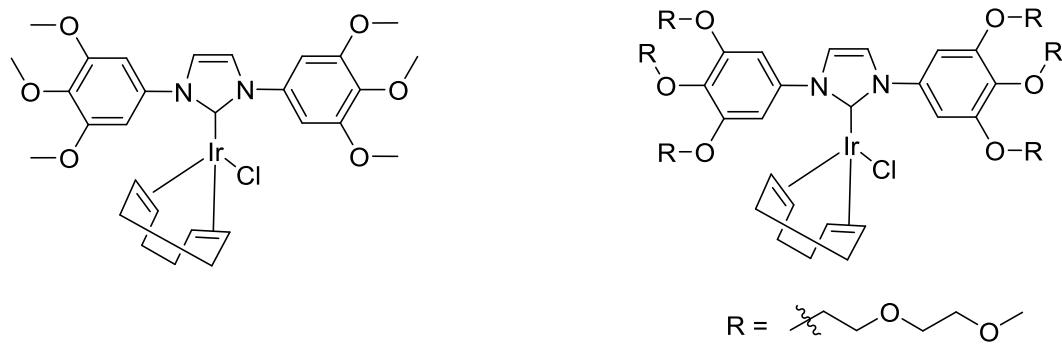


Figure 25. Chemical structures of the pre-catalysts $[\text{Ir}(\text{Itome})(\text{COD})\text{Cl}]$ and $[\text{Ir}(\text{IDEG})(\text{COD})\text{Cl}]$.

Chekmenev and co-workers have similarly modified the NHC, but in an asymmetric fashion [93]. The PEGylated iridium catalyst as shown in Figure 26 was synthesised. This catalyst proved to have excellent solubility in water. After polarisation transfer at low-field (un-optimised, $\sim 11 \pm 5$ mT), the

¹H NMR signals of pyridine gave a 16-fold enhancement in *d*₆-ethanol. The addition of D₂O to *d*₄-methanol solutions had a negative effect on the enhancement and as such had lower enhancement values than those in *d*₆-ethanol (20% D₂O / 80% *d*₄-methanol yielded only a 6-fold enhancement). Higher amounts of D₂O led to negligible or zero enhancement being detected. Increasing the pressure of *parahydrogen* to 60 psi led to better polarisation efficiency as a 110-fold enhancement for pyridine was obtained. However, when a 63% D₂O / 37% *d*₆-ethanol solution was used, no SABRE enhancement was observed. Based on this observation, a second non-PEGylated pre-catalyst was synthesised which utilised 1,2-dihydroxy-3,7-cyclooctadiene (CODDA) instead of COD to give [Ir(IMes)(CODDA)Cl]. Although not as soluble as the PEGylated complex, the CODDA pre-catalyst was water soluble to some extent (0.3 mg ml⁻¹). Once activated, the catalytic form is [Ir(IMes)(py)₃(H)₂]⁺, and under an atmosphere of *parahydrogen* for 30 s in a field of ~10 mT an enhancement of 25-fold for the *ortho* ¹H NMR signal of pyridine. This enhancement was reproducible in subsequent repeats of the experiment. Notably, an aqueous suspension of [Ir(IMes)(COD)Cl] yielded no SABRE enhancement following activation. Thus, the use of CODDA pre-catalyst circumvents the need to dissolve in organic solvent, dry and then reconstitute in D₂O [51, 91], whilst maintaining the same level of enhancement.

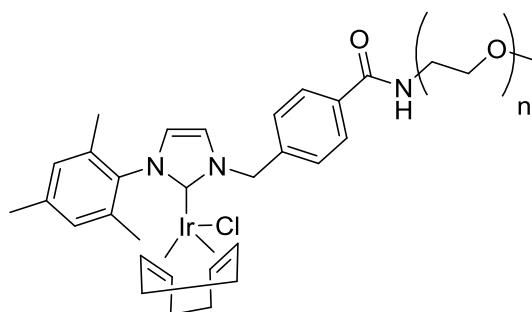


Figure 26. Pre-catalyst synthesised by Chekmenev and co-workers (*n* = 12) [93].

Vitamin B3 (nicotinic acid) has been polarised by SABRE and the signal enhancement of the four aromatic proton nuclei recorded as a function of pH [94]. In this work, nicotinic acid was polarised along with its conjugate base produced by adding increasing amounts of caesium carbonate to the solution. Compared with the uncharged form, production of the conjugate base led to the greater signal enhancement; at pH 8 the total signal enhancement over all four proton sites was *ca.* 400-fold whereas at pH 12 the signal enhancement increased to over 800-fold. Interestingly, the *T*_{1s} of the uncharged and conjugate base forms have similar *T*_{1s} in the presence of transfer catalyst, aside from the H-2 signal which is 7.3 s longer for the conjugate base and thus, closely mirrors that of free nicotinic acid (*T*₁ = 18.4 s). The observation of a single hydride resonance in the ¹H NMR at high pH demonstrated that the conjugate base dominates and as such, no speciation of the catalytic species occurs.

Heterogeneous SABRE

With a future view of being able to separate the catalyst from a hyperpolarised bolus (thus maximising the longevity of the hyperpolarised state), heterogeneous catalysts for use with SABRE have been designed and synthesised. The use of heterogeneous catalysts can circumvent problems associated with the isolation of the catalyst and the production of by-products associated with catalyst activation (for example cyclooctane, the hydrogenation product of COD). Tethering of [Ir(IMes)(COD)][PF₆] on to polymer microbeads, which possess pyridyl surface groups) have been accomplished by Shi *et al.* [95]. Successful immobilization of the iridium complex on to the polymer microbeads was supported by AAS, IR and MS. AAS revealed that 4% of the total heterogeneous

SABRE catalyst was iridium complex. A hyperpolarised ^1H NMR spectrum, following polarisation transfer at 100 G, was obtained using the heterogeneous SABRE catalyst which was qualitatively similar to that obtained using the homogeneous catalyst. Only signals for free pyridine were observed. Signals for those of the catalyst and free pyridine were not observed as the microbeads settled quickly to the bottom of the sample tube once shaking of the tube to enact polarisation transfer had ceased. The *ortho*-, *para*- and *meta*-proton nuclei had their respective ^1H NMR signals enhanced by 5.2-, 4.1- and 2.7-fold respectively. All of these signals were observed to be in emission. Under the same conditions, the homogenous catalyst produced values of *ca.* 70-fold. Isolation of the supernatant from a sample that contained the heterogeneous catalyst was evaluated for SABRE activity, primarily to see if catalyst leaching had occurred. Following polarisation transfer at 100 G, no enhancement of pyridine was observed. AAS (Atomic Absorption Spectroscopy), MS (Mass Spectrometry) and long ^1H NMR acquisitions correlate with this observation as the presence of leached catalyst was not detected. Thus, the heterogeneous support employed was robust and did not leach catalyst; the enhancements observed were purely associated with polarisation transfer derived from heterogeneous, rather than homogenous, catalysis. This also infers that the pyridyl linker of the microbeads is bound to the catalyst *trans* to IMes, and not *trans* to hydride, otherwise catalyst would dissociate from the support (and *vice versa*) during polarisation transfer.

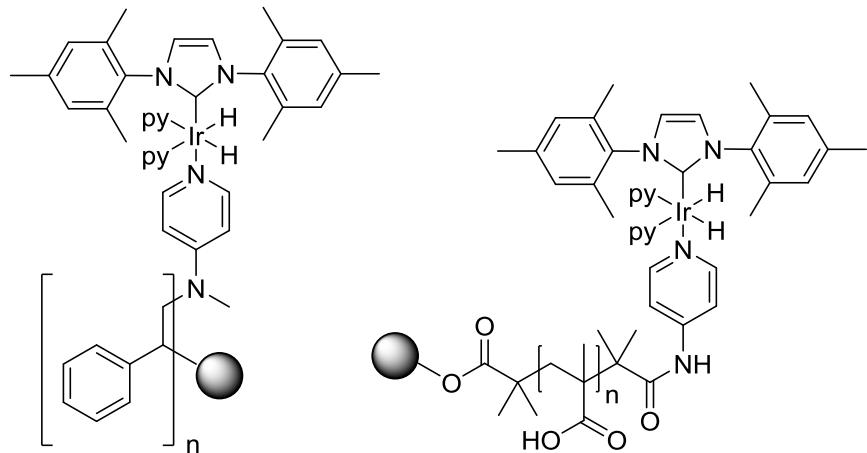


Figure 27. Heterogeneous SABRE catalysts, which have been either attached to polymer microbeads (left [95]) or TiO_2 / PMAA nanoparticles (right [96]).

In a further study, the SABRE catalyst was heterogeneously supported on to a TiO_2 / PMAA (poly(methacrylic acid)) core / shell nanoparticle or PVP (polyvinylpyridine) “comb” polymers to produce two new heterogeneous SABRE catalysts [96]. The PMAA shell of the TiO_2 / PMAA nanoparticles was functionalised with 4-aminopyridine. Both catalysts were characterised by AAS, NMR and DLS (Dynamic Light Scattering). Based on AAS results, 25% of the total nano-SABRE catalyst particles by weight is the iridium complex. Analysis of filtrates by AAS isolated from the final washing step of the production of the heterogeneously supported catalyst and after catalyst activation in the presence of H_2 , did not detect the presence of iridium. The catalysts are again robust under the conditions in which they are employed. The PVP “comb” polymer heterogeneous SABRE catalyst enhanced the ^1H NMR signals of the *ortho*-, *para*- and *meta*-sites of pyridine by 7-, 6- and 3-fold respectively. Employment of the TiO_2 / PMAA core / shell nanoparticle-based heterogeneous SABRE catalyst resulted in the pyridine ^1H NMR signals being enhanced by 18-, 17 and 7-fold respectively for the *ortho*-, *para*- and *meta*-positions. These enhancements were improved following 300 s of *ex situ* parahydrogen bubbling; the values obtained as a result were increased to 26-, 39- and 11-fold respectively (*ca.* 0.13% ^1H polarisation). The feasibility of recovering and recycling the supported

catalysts was also demonstrated. The enhancements obtained were slightly reduced, but still comparable to those obtained prior to the recovery process.

Another study has looked at deactivating the SABRE catalyst, using bipyridine (bpy), in order to improve the longevity of the hyperpolarised signal [97]. In the presence of the catalyst, $[\text{Ir}(\text{IMes})(\text{CODCl})]$, the T_{1s} of nicotinamide were substantially reduced compared to solely nicotinamide. These are shown in Table 3. The addition of an equimolar amount of bpy relative to nicotinamide led to the production of a new complex, $[\text{Ir}(\text{IMes})(\text{bpy})(\text{H})_2(\text{nicotinamide})]^+$, in which bpy is *trans* to both hydride ligands and nicotinamide is *trans* to IMes. Subsequent re-polarisation of the same solution, that now contains $[\text{Ir}(\text{IMes})(\text{bpy})(\text{H})_2(\text{nicotinamide})]^+$ and $[\text{Ir}(\text{IMes})(\text{H})_2(\text{nicotinamide})_3]^+$ led to T_{1s} of nicotinamide which were comparable to that of free nicotinamide. Notably, after 3 T_{1s} (19.5 s), H-2 of nicotinamide retains only 5% of its hyperpolarised signal, whereas after the addition of bpy, it takes 136 s to reach the same level. Re-polarisation of this sample resulted in a significant drop in the enhancement observed; polarisation transfer efficiency has been reduced considerably. Indeed, due to the dissociation rate of nicotinamide from $[\text{Ir}(\text{IMes})(\text{H})_2(\text{nicotinamide})_3]^+$ being $6.7 \pm 0.1 \text{ s}^{-1}$, the minimum time for 97% conversion to $[\text{Ir}(\text{IMes})(\text{bpy})(\text{H})_2(\text{nicotinamide})]^+$ is *ca.* 1.2 s and so rapid quenching is achievable. Use of UV-light to remove bpy and subsequently enable the solution to be utilised for polarisation studies again, proved unsuccessful. Use of a more rigid quenching agent, 1,10-phenanthroline also resulted in elongated times that the hyperpolarised signal persisted for. Using a series of low-tip angle pulses, the ^1H NMR signals of nicotinamide took *ca.* 170 s to return to thermal equilibrium. When nicotinamide was exchanged for pyridine, and the catalyst quenched with 1,10-phenanthroline, the hyperpolarised signal persisted for 94 s. Although this value is smaller than for nicotinamide, it must be noted that only 5 equivalents of pyridine relative to catalyst was used (instead of five in the case of nicotinamide) and the initial signal intensity across all the proton sites was *ca.* 300-fold higher than that observed for nicotinamide.

Table 3. Reported T_1 data of nicotinamide in d_4 -methanol under a H_2 atmosphere with and without $[\text{Ir}(\text{IMes})(\text{H})_2(\text{nicotinamide})_3]^+$ and $[\text{Ir}(\text{IMes})(\text{H})_2(\text{nicotinamide})_3]^+$ at 9.4 T.

Proton label	T_1 / s (free form)	T_1 / s (in the presence of $[\text{Ir}(\text{IMes})(\text{H})_2(\text{nicotinamide})_3]^+$)	T_1 / s (in the presence of $[\text{Ir}(\text{IMes})(\text{bpy})(\text{H})_2(\text{nicotinamide})]^+$)
H-2	43.1	7.0	38.0
H-4	11.3	6.1	11.5
H-5	13.5	3.7	14.0
H-6	6.6	4.3	6.8

Deuteration of analyte molecules can improve the signal enhancement observed, provide insight into the polarisation transfer mechanism and increase the longevity (T_1) of the signal. This is despite the preparation of these materials sometimes being synthetically challenging. The employment of a deuteration strategy has been explored for nicotinamide and methyl nicotinate to maximise the longevity of the hyperpolarised signal [98]. For four equivalents of nicotinamide, in the presence of $[\text{Ir}(\text{IMes})(\text{CODCl})]$, the four aromatic ^1H sites are enhanced by *ca.* 2050-fold over in d_4 -methanol. The T_{1s} range from 3.7 – 6.3 s, but in ethanol they are 2.4 – 4.2 s and in D_2O they are 6.5 – 10.7 s, all of which are in the presence of $[\text{Ir}(\text{IMes})(\text{CODCl})]$ and under a hydrogen atmosphere. The longest T_{1s} were obtained when either the 5,6-, 2,5- or the 4,6-positions were deuterated. The biggest polarisation, however, was obtained for 4,6- d_2 -nicotianamide; the protons were polarised to 4.1% and the T_{1s} were 26.8 and 47.9 s in d_4 -methanol for positions 2 and 5 respectively. In ethanol the T_{1s} reduced to 8.2 (H-2) and 17.5 (H-5) s whilst in D_2O they were 37.1 (H-2) and 54.0 (H-5) s. Having

established that the $^5J_{\text{HH}}$ coupling present between the two protons in 4,6- d_2 -nicotianamide is sufficient for propagating polarisation (as exemplified by the polarisation levels observed on both sites), methyl nicotinate was then investigated. The methyl group of methyl nicotinate was deuterated as it was hypothesised the amide protons were aiding the relaxation process. The polarisation observed for d_3 -methyl-4,6- d_2 -nicotinate rose to 8.9%. The T_{1s} in d_4 -methanol were commensurate with those of 4,6- d_2 -nicotianamide, whilst the average T_1 in D₂O increased. Interestingly, 4,6- d_2 -nicotinate possessed better polarisation values than d_3 -methyl-4,6- d_2 -nicotinate, in addition to retaining similar, albeit slightly reduced T_1 values in all three solvents. This substrate was, therefore, chosen for optimisation. Increasing the *parahydrogen* pressure to 4.0 and 5.5 bar resulted in the polarisation of H-2 increasing to 26.4% and 28.5% respectively. For this study, [Ir(d_{22} -IMes)(COD)Cl] was used. Use of [Ir(d_{22} -IMes)(COD)Cl] increased further the T_{1s} of the two ¹H nuclei of 4,6- d_2 -nicotianamide as it minimised polarisation loss due to relaxation processes. The polarisation value of H-2 was increased to 50% (45% average over both sites) by employing a sacrificial co-ligand, methyl-2,4,5,6- d_4 -nicotinate.

Imaging Applications of SABRE

The production of hyperpolarised molecules that possess much improved signal intensities compared to their thermal polarisations, has resulted in investigations to probe their use as contrast agents for MRI. Analytes possessing ¹H, ¹³C, ¹⁵N or ³¹P nuclei have been hyperpolarised to produce MR images using a phantom-based setup. Literature examples relating to this are presented in the following sections. At present, *in vivo* images have not been reported, but the following studies provide a strong foundation for this to occur in the near future.

¹H

A field-cycling method has been detailed to obtain images at earth's magnetic field [99]. This work made use of a continuous-flow, high-throughput, high-pressure *parahydrogen* converter which produces a flow rate of four litres per minute of ~98% *para*-enriched hydrogen gas [15]. Following continuous polarisation transfer, the theory of which has been described [100], in a magnetic field of 6.5 mT for 4 s, which took place within the MR imager, the field was then switched off to facilitate signal excitation, spatial encoding and read-out at earth's magnetic field. Imaging was limited to acquiring one line of k space at a time (repetition time 8 s) and thus represents an area for improvement so that images are obtained quicker. An image of a hyperpolarised solution containing 80 mM of analyte with a field of view 64 mm² was produced with sufficient signal-to-noise to give an interpolated resolution of 1 mm². Total acquisition time was 4.16 minutes. An image of a solution of water (56 M) that was pre-polarised at 20 mT took 2.16 hours to acquire to give roughly the same signal-to-noise. Therefore, the hyperpolarised image represents a considerable time saving. With a view towards possible *in vivo* use, continuous hyperpolarisation of pyridine in ethanol using [Ir(IMes)(COD)Cl] with and without the addition of fresh human blood to the solution was probed. The polarisation values obtained are much reduced compared to a pure methanol solution and decrease further upon the addition of human blood. After obtaining one hundred scans (with a repetition time of 10 s), the solution had coagulated. Although the enhancements obtained were lower than those in methanol, the use of continuous hyperpolarisation offers an attractive route to *in vivo* experiments.

This work was subsequently extended to an aqueous system to polarise nicotinamide continuously at low magnetic field [87]. The method developed by Truong *et al.* was utilised to obtain aqueous SABRE solutions [51]. At 5.6 mT, a hyperpolarised NMR signal with a signal-to-noise of 56 was detected and maintained for more than 1000 s. The polarisation level, although only 1.6x10⁻⁴, does

benefit from the enhancement being continuously available, and therefore, adequate signal-to-noise can be easily obtained through repetitive acquisitions. A ^1H MR image was obtained of continuously hyperpolarised nicotinamide (70 mM) in 5:20 minutes. This gave a signal-to-noise of 13.4 compared with water (55 M) which had a signal-to-noise of 4.7. The relatively long time to collect the image was due to the requirement to field-cycle for hyperpolarisation (5.9 mT) and the measurement field (52 μT) for every k -space line. Without this requirement, it was estimated that the image could be acquired in *ca.* one minute.

Low-field MR images have been acquired of pyridine following polarisation under SABRE conditions [101]. Images were acquired using two different approaches; *in situ* and *ex situ*. The *in situ* approach utilises the same magnetic field for measurement and polarisation transfer (5.75 mT or 47.5 mT). *Ex situ* utilised a magnetic field of 5.75 mT for polarisation transfer and then a magnetic field of 47.5 mT for detection. The high-field SABRE effect is required to generate *in situ* images at 47.5 mT. *In situ* 2D MR images acquired at 47.5 mT possessed $0.75 \times 0.75 \text{ mm}^2$ in-plane spatial and 4.5 min temporal resolution. Due to the higher polarisation yielded using the *ex situ* method, better in plane spatial ($0.125 \times 0.125 \text{ mm}^2$) and temporal resolution ($\leq 7.7 \text{ s}$) were obtained. Both images have good spatial resolution, despite *ex situ* experiments producing signals that are two orders of magnitude larger than *in situ* experiments. The *in situ* spatial resolution is also aided by being able to use a larger excitation RF pulse tipping angle and signal averaging, due to constantly renewed polarisation that builds up quickly (time constant for build-up = $7.4 \pm 0.3 \text{ s}$). Thus, temporal resolution can be sacrificed at the expense of improving spatial resolution greatly. Compared with a method that utilises field cycling to obtain hyperpolarised images [99], *in situ* detection enables a line of k -space to be acquired every 0.2 s instead of 8 s and gives better spatial resolution than demonstrated previously.

The effect of deuteration on the analyte on the intensity of the MR image has been demonstrated for pyridine [102]. Compared to pyridine, which gave a 169-fold contrast enhancement relative to a water internal standard, 3,4,5- d_3 -pyridine produced a 807-fold enhancement ($P = 2.5\%$). The decrease in the number of proton sites to accept polarisation increases the signal enhancement in the same fashion as that observed in ^1H NMR SABRE studies.

The effect of pH on the resulting signal intensity observed in MR images has been exemplified for vitamin B3 [94]. Samples consisted of 5 mM of $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ and 20 equivalents of vitamin B3 in d_4 -methanol. An increasing amount of caesium carbonate (up to 17-fold excess) was added to the sample and MR images acquired at pHs 3.6, 4.6, 5.1, 5.7, 5.9 and 6.9 with a voxel of $3 \times 3 \times 10 \text{ mm}$. Just like the ^1H NMR measurements, these images increase in signal intensity with increasing pH. However, it was noted that at physiological pH and above, the signal-to-noise ratio plateaus before falling rapidly beyond pH 8. This was thought to be due to the possible transverse relaxation rate effects caused by the methoxide ion. A ^{13}C image of ^{13}C -labelled vitamin B3 (^{13}C located on the quaternary of the carboxylic acid) was acquired; the signal-to-noise of the uncharged vitamin B3 did not yield signals above the background noise whereas those of the conjugate base did.

^{13}C

^{13}C is naturally only present at the 1% level. This makes ^{13}C labelled racers for MRI an exciting prospect. Thus far, however, there has been limited reports of ^{13}C -labelled materials towards this purpose in conjunction with SABRE. The synthesis of ^{13}C -labelled nicotinamide (^{13}C located in amide) represents one example and this has enabled the collection of SABRE hyperpolarised ^{13}C images [102]. A single scan image of ^{13}C -labelled nicotinamide gave a polarisation value of 0.03%. Although

this value is low, to obtain a comparative image, a total scan time of 18 hours 20 minutes was required (1100 averages).

^{15}N

The ease of obtaining a hyperpolarised ^{15}N state using SABRE-SHEATH has motivated a ^{15}N MRI study [73]. ^{15}N labelled pyridine ($\sim 24 \mu\text{mol}$) was hyperpolarised to deliver ca. 1% polarisation following a long 30 s transfer from the magnetic shield in which polarisation was conducted to the 4.7 T MRI scanner in which the measurement was performed. An axial 2D gradient echo ^{15}N MR image with $2 \times 2 \text{ mm}^2$ spatial resolution was acquired and this image showed hyperpolarised liquid within a 6.6 mm internal diameter high-pressure HPLC (High Performance Liquid Chromatography) column. This experimental setup results in the internal diameter of the column being imaged using less than four pixels. The authors note that the long-lived polarisation that is generated on the ^{15}N nuclei could be partially transferred to ^1H nuclei to obtain ^1H MR images. The partially depleted polarised state could be utilised many times to produce subsequent ^1H MR images. Thus, this may offer a route to track these molecules *in vivo* or in metabolic processes.

^{31}P

In addition to ^1H , ^{13}C and ^{15}N MR images, images utilising the ^{31}P nuclei in a sample have also been acquired to demonstrate the application of SABRE. ^{31}P images of a 6 mM solution of $[(\text{PPh}_3)_3\text{Ir}(\text{H})_2\text{Cl}]$ in toluene were acquired thermally and hyperpolarised at 80°C in a μ -magnetic shield [36]. The thermal image produced does not produce a resolved image in which the cross-section of the 15 mm sample tube is evident. This is despite the signal being comprised of 128 transients. However, the corresponding hyperpolarised single scan ^{31}P image clearly resolves the sample tube. The hyperpolarised scan took less than one second to acquire whereas the thermal image took hours. There is quite clearly a resolution boost as well as time savings whilst employing SABRE.

The aforementioned study utilises toluene as the solvent. This is not biocompatible with a living organism. Duckett and co-workers reported RARE images of the deuterated analyte shown in Figure 28 in deuterated methanol [103]. Despite this still not being a biologically relevant medium, it does represent a move towards acquiring ^{31}P images in aqueous media. An image of the hyperpolarised sample was acquired in 500 ms, and relative to the 2048 scan thermal image which took twelve hours of measurement time. The hyperpolarised image showed >8 -fold improvement in signal-to-noise relative to the thermal image. The undeuterated analogue showed a >4 -fold improvement in signal-to-noise. Conducting the experiment in a horizontal scanner led to a reduction in the SNR for the deuterated *bis*-3,5-substituted derivative as there was greater residual motion of the sample and a longer period of time to transfer the sample to the magnet (ca. 5 s). As the T_1 of the ^{31}P nuclei in this substrate are only 6 s, transfer from the polarisation transfer field to the measurement field must be as rapid as possible to ensure integrity of the polarised state. Thus, to utilise these types of ^{31}P containing substrates, the T_1 needs to be substantially longer for biomedical use.

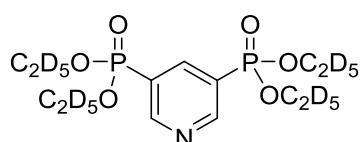


Figure 28. Chemical structure of the deuterated analogue utilised by Burns *et al.* to obtain hyperpolarised ^{31}P images [103].

Generation and utilisation of long-lived states by SABRE

The relatively short T_1 s that some nuclei possess, such as ^1H , limits the usefulness of the resulting hyperpolarised molecules. The section on polarisation transfer to ^{15}N highlights this is not a problem

for ^{15}N . To offset this shortcoming, without having to resort to deuteration strategies, utilisation of long-lived states have been investigated. 2-aminothiazole, atz, is a molecule that readily polarised by SABRE yielding enhancements of 386 ± 20 and 432 ± 38 fold for the H-4 signal for 17- and 2-fold equivalents relative to $[\text{Ir}(\text{IMes})(\text{atz})_3(\text{H}_2)]^+$ (atz binds through the thiazole ring nitrogen) [104]. The respective T_{1s} for H-4 were 12.2 ± 0.5 and 16.2 ± 0.6 s respectively at the same level of equivalence. The use of a 1 s 1 kHz low power spin-lock, following polarisation transfer by SABRE at low-magnetic field, created a dramatic increase in the singlet magnetisation. Increasing the spin-lock time to one minute still led to the observation of significant hyperpolarised singlet derived magnetisation. The spin-lock, therefore, acts to protect the created singlet state.

This approach was extended to deuterated nicotinamide and pyrazine derivatives [105]. Singlet lifetimes for all three derivatives were in the order of *ca.* 40 – 50 s, which were significantly longer than their corresponding T_1 values. The longest was seen for methyl-3-*d*-pyrazine-2-carboxylate, which also possessed the smallest chemical shift difference between its two proton nuclei (34.4 Hz). The shortest T_{LLS} (LLS = long-lived singlet state) was observed for 2,4-*d*₂-nicotinamide; the two proton nuclei are separated by a chemical shift difference of 457.0 Hz. No significant variation in T_{LLS} was observed relative to substrate equivalence. This contrasts to signal enhancement which decreases with increasing substrate excess. The maximal amount of singlet order was produced following polarisation transfer that generated the largest signal SABRE derived enhancement.

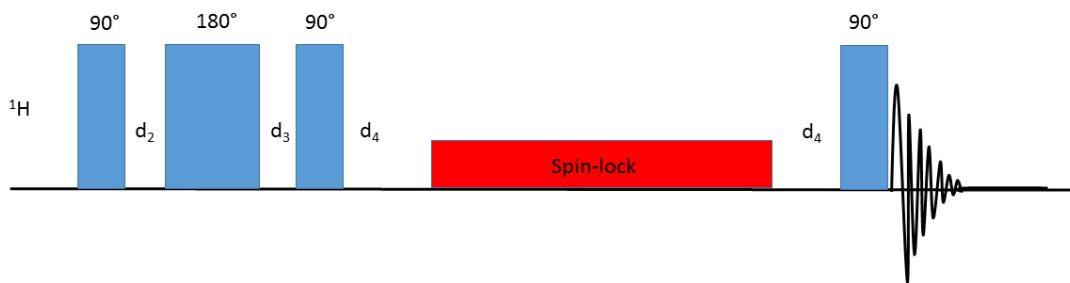


Figure 29. Pulse sequence used to create, store and read-out singlet states. $d_2 = 1/4J$, $d_3 = 1/4J + 1/2\Delta\nu$ and $d_4 = 1/4\Delta\nu$

The importance of substrate choice in generating the longest possible T_{LLS} times was highlighted in a further study on pyridazine derivatives [106]. Despite the steric bulk surrounding the two nitrogen atoms of 2-methyl-6-*d*₃-methylpyridazine, an enhancement of 650-fold was observed. In addition, there only exists a $^5J_{\text{HH}}$ coupling between a *parahydrogen* derived hydride and those of the proton nuclei of 2-methyl-6-*d*₃-methylpyridazine to propagate polarisation. Of the series, 2-methyl-6-*d*₃-methylpyridazine possessed the smallest chemical shift difference between its two proton nuclei (0.5 Hz). In a magnetic field of 9.4 T, T_{LLS} was 32 ± 1 s, whereas at low-field (*ca.* 10 mT), T_{LLS} improved to 255 ± 23 s. Hyperpolarised signal was still visible 15 minutes after its creation. For 1-*d*₃-methyl-6-*d*-pyridazine, a chemical shift difference of 1 Hz exists between the proton nuclei. Due to the decrease in steric hindrance, the enhancement observed under SABRE increases to 2040 with T_{LLS} being 113 ± 4 s at high-field with spin-locking. This material's hyperpolarised signal was still detectable 10 minutes after generation. In general, about 66% magnetisation to singlet conversion was achieved. The long periods for which the magnetisation can be stored is useful for a biological application as it allows the hyperpolarised sample to be transported to the final measurement location without any signal loss.

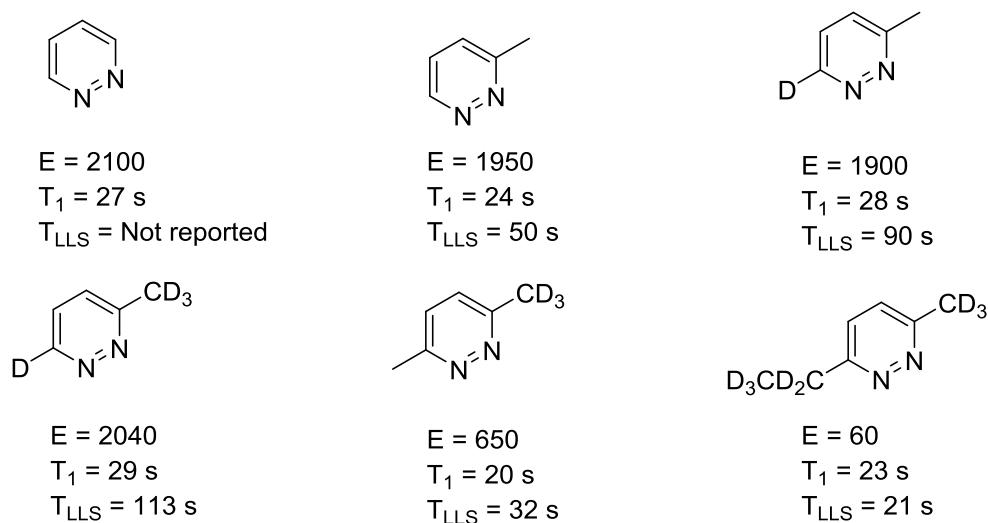


Figure 30. Pyridazine substrates and their enhancements (E), T_1 values and T_{LLS} at high-field with spin-locking.

SABRE-SHEATH has also been used to create long-lived states in two pyridyl acetylene derivatives, one of which was symmetric, the other asymmetric [107]. The acetylene component was ^{13}C -labelled and compared to analogous non-labelled derivatives. Employment of SABRE-SHEATH led to enhancements of 37- to 166-fold. Derivatives that were not ^{13}C -labelled, were enhanced by 2-3 times more than ^{13}C -enriched derivatives. This may be due to faster T_1 relaxation in the $^{13}\text{C}_2$ pairs as opposed to isolated ^{13}C spins. $[\text{Ir}(\text{IMes})(\text{COD})\text{Cl}]$ was used as the pre-catalyst for these experiments. Interestingly, no hyperpolarisation of the ring carbons resulted and hydrogenation of the alkyne bond was observed. Both of these observations would suggest that binding was occurring predominantly through the alkyne bond rather the pyridyl motif. However, diphenyl- $^{13}\text{C}_2$ -acetylene did not polarise and ^1H NMR spectra revealed strong SABRE enhancements for the pyridyl protons. More likely is that the T_{1s} of carbon nuclei with protons directly attached are short (~ 5 s) and, therefore, prevent build-up and detection. Conversely, the T_{1s} of the bridge and acetylene carbon nuclei are 11 and 12 s respectively at low-field. Magnetisation and singlet order are produced in the ^{13}C -labelled derivatives. The singlet order has a lifetime of ~ 2 minutes. Read-out of the singlet order in the asymmetric derivative occurs readily as the singlet state is no longer an eigenstate (chemical shift difference ($\Delta\nu_c \sim 770$ Hz) is larger than the J_{cc} coupling (185 Hz)). This is not true for the symmetric derivative and as such, specialised pulse sequences would be needed to convert the singlet order in to detectable magnetisation.



Figure 31. The symmetrical and asymmetrical acetylene derivatives that were investigated by Theis and co-workers [107].

Summary

This work has been concerned with detailing the hyperpolarisation technique SABRE and its applications in NMR and MRI. The chemical structure of the pre-catalyst was detailed and the importance of the inclusion of a carbene ligand compared to a phosphine ligand highlighted in terms of improving the enhancements obtained. The physical attributes that the catalyst needs to possess were covered and this demonstrated the complex interplay of a number of constraints such as *parahydrogen* pressure, solvent selection, amount of substrate, the magnetic field in which polarisation is conducted, type of substrate to be polarised and the temperature at which polarisation transfer is conducted. Analysis of parameters such as PAAP, and the use of DFT, for

example, were detailed to explain further how insight is being gleaned about the polarisation transfer process. Exchange rates of the substrate and the hydride ligands were discussed to bring to the reader's attention the reversible nature of the process and how they can be tailored to improve polarisation efficiency. Works on the utilisation of other group 9 congeners as the central metal of SABRE catalysts and the utilisation of bi- and tri-dentate ligands in catalyst construction, were also presented.

The applications of SABRE were highlighted, using examples from the literature. These mainly focused on detecting species at low concentrations or magnetic field that would be difficult using conventional NMR. The hyperpolarisation of ^1H , ^{13}C , ^{15}N , ^{19}F , ^{31}P , ^{29}Si and ^{119}Sn nuclei were the subject of these various studies. Synthetic strategies, whereby molecules are synthesised to maximise polarisation transfer or deuterated to minimise the number of polarisation acceptors, were stressed as methods to improve signal intensities. The use of co-ligands were emphasised as a further way to increase the observable enhancement. The collection of 2D NMR spectra, such as COSY and HMBC, were discussed to demonstrate how SABRE improves the quality of these spectra, as well as reducing the time required to acquire them. Utilisation of ultra-fast techniques was another area highlighted that employed SABRE whereby 2D spectra have been acquired in a single scan.

Special emphasis was placed upon SABRE-SHEATH based experiments, as due to the wealth of research conducted recently, this is a readily apparent hot-subtopic of SABRE. The scope of the molecules polarisable by this technique were presented, and the rationale behind their relative enhancements provided. In particular, the employment of SABRE-SHEATH to polarise non-*N*-heterocycles represents an exciting breakthrough for SABRE and further diversifies what this technique can be applied to.

High-field SABRE was also presented to show that SABRE is not limited to low-field derived polarisation transfer. The improvements of this technique were detailed, by showcasing relevant pulse sequences. In addition, the employment of the technique for the collection of 1D and 2D NMR spectra of complex mixtures, and for the chemosensing of natural extracts, was exemplified.

The hyperpolarisation of biologically relevant analytes was detailed, as was the conducting of SABRE based experiments in aqueous systems. The motivation behind these studies has been mainly to demonstrate that SABRE can be translated from a NMR-based approach to MRI, as there is currently considerable interest in the production of non-gadolinium based contrast agents. Reports detailing the collection of ^1H , ^{13}C , ^{15}N and ^{31}P MR images of phantoms were exemplified and the merits of the acquired images highlighted. Although a SABRE-based *in vivo* image is yet to be disseminated in the literature, the works detailing imaged phantoms provides a strong foundation for this to be achieved.

The work concluded with a section on long-lived states. This area presented the current research on elongating the T_{1s} of substrates by accessing these substrates. As the T_1 is an effective measure of the longevity of polarisation, maximising this results in a longer time-frame to successfully deploy the hyperpolarised material prior to magnetisation read-out.

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